

Aliphatic Halide-Carbonyl Condensations by Means of Sodium

By Edgar A. Cadwallader, Abraham Fookson, Thomas W. Mears, and Frank L. Howard

As a part of an investigation of the synthesis of highly branched aliphatic hydrocarbons that is being conducted at the National Bureau of Standards for the National Advisory Committee for Aeronautics, the Navy Bureau of Aeronautics, and the Army Air Forces, several compounds have been prepared by interaction of alkyl halides and various carbonyl compounds in the presence of sodium. This reaction makes possible the synthesis of certain highly branched compounds not easily obtainable by other means.

I. Introduction

The use of the Grignard reaction for synthesis of highly branched compounds is limited by undesirable side reactions, which become progressively more pronounced as the degree of branching of the reactants is increased. These side reactions involve the reduction and enolization of the adduct rather than the desired addition of the alkyl group.

The recent work of Bartlett and Schneider [1]¹ and previous work by Morton and Stevens [13] indicate that the reaction between an alkyl halide and an ester or ketone under the influence of sodium may be used to prepare certain compounds more easily than by the Grignard reaction, or may be used in some cases in which the Grignard reaction fails completely as a synthetic tool. In this paper, are described the results of a few reactions that were designed to give information concerning the scope of the reaction and the effect of various solvents and temperatures of reaction. The work of Bartlett and Schneider has been extended to include a few additional reactants.

II. Experimental

The apparatus used in making small-scale runs was the conventional three-neck flask of appropriate size, fitted with a stirrer, reflux condenser, and thermometer. A separatory funnel was placed at the top of the condenser. Heating and

cooling of the reaction mixture was effected by application of hot or cold oil baths. In some cases, particularly with low-boiling solvents, both small necks of the flask were fitted with reflux condensers with the elimination of the thermometer.

Sodium sand was prepared batch-wise by heating a weighed amount of sodium in a purified light mineral oil until melted, and then shaking the liquid sodium vigorously in a closed Erlenmeyer flask until the desired state of subdivision was obtained. After the particles of sodium had cooled sufficiently, they were washed by decantation with three or four portions of the solvent to be used, transferred to the reaction vessel, and immediately covered with solvent.

A few of the first reactions tried in this series were conducted in *n*-pentane (boiling range 34° to 38° C), but later ones were conducted in isooctane² (boiling point, 99° C). It was found that the vigorous reactions were more easily controlled when the higher-boiling solvent was used. Also the use of isooctane as a solvent facilitated the use of higher temperatures for the reactions.

Bartlett and Schneider carried out their reactions by adding all of the alkyl halide to the sodium at once and then adding the carbonyl compound over a period of time. In attempting to use this

¹ This is the commonly used name for 2,2,4-trimethylpentane, which is a primary standard for octane number measurements. The commercially obtained material is exceptionally pure and serves admirably as an inert solvent for reactions of this type. The common name is used throughout this paper.

² Figures in brackets indicate the literature references at the end of this paper.

technique in our applications of the reaction, we often found that the reaction was so vigorous that the sodium was completely consumed before all of the carbonyl compound could be added. Apparently in these cases there is a self-sustaining reaction between the sodium and the alkyl halide that needs only a starting impetus given by the carbonyl compound. In order to minimize this effect only a small portion of the alkyl halide was added at the start of the reaction, and the remainder was mixed with the carbonyl compound for gradual addition during the reaction.

After the reaction had been carried to completion, the mixture was decomposed either by pouring it into ice water or by direct addition of water to the mixture in the flask. If any sodium was visible in the flask, the treatment with water was preceded by the addition of small amounts of ethyl alcohol. After complete hydrolysis, the two layers were separated and the organic layer washed with several portions of water. The organic substance was finally dried over a suitable drying agent. In those cases where the products of reaction were well known, only those physical constants that were deemed necessary to identify the product were determined.

Some of these reactions have been used for large-scale synthesis; the equipment used in these cases has been described previously [8].

Methyl isobutyrate, *t*-butyl chloride, methyl *n*-butyrate, isobutyl chloride, *n*-propyl chloride, and *n*-butylchloride were all made by standard procedures. Ethyl trimethylacetate and hexamethylacetone (2,2,4-tetramethyl-3-pentanone) were prepared by methods previously described [8]. Pentamethylacetone (2,2,4-trimethyl-3-pentanone) was prepared by a large-scale run of the reaction between ethyl trimethylacetate and isopropyl chloride by the method described herein. Pinacolone (3,3-dimethyl-2-butanone) [8] was prepared by the oxidation of 3,4,4-trimethyl-2-pentene [8] with sulfuric acid-dichromate solution in a manner analogous to that used previously for the preparation of 4,4-dimethyl-2-pentanone [8]. The other materials were obtained commercially.

Only constant-boiling materials that were collected from distillation in a 180 by 2.5-cm Whitmore and Lux [16] column were used in the synthesis. Several of these columns were used; some were packed with glass helices, the others were packed with stainless steel helices, but all

were equipped with the type of distilling head recently described [7].

Solvents used for the reactions were removed by distillation through these columns; the residues were distilled (generally at reduced pressures) in a Podbielniak Hyper-cal column or in a Claisen flask fitted with a short fractionating column.

Yields were calculated on the basis of the amount of carbonyl compound that was added to the reaction. Table 1 summarizes data obtained on the several products, together with results reported in the literature.

1. Diisopropyl Ketone and *n*-Propyl Chloride

Sodium sand (69 g, 3 g atoms) was added to 1 liter of isooctane in the reaction vessel. The contents was heated to 80° C, and a solution of 114 g of diisopropyl ketone (1 mole) and 157 g of *n*-propyl chloride (2 moles) was gradually added. After the reaction began, the temperature rose to 95° C, which caused considerable refluxing. The characteristic¹ greenish-black color appeared, and addition of the ketone-halide solution was continued at such a rate as to insure vigorous refluxing. Addition time was approximately 2 hours. After standing overnight the reaction mixture was hydrolyzed, and the organic layer was washed and dried. Isooctane and unreacted *n*-propyl chloride were removed by distillation, and the residue was distilled at reduced pressure (Claisen flask). There was obtained 40 g of recovered diisopropyl ketone, and 51 g of 2-methyl-3-isopropyl-3-hexanol (32% yield). This carbinol has been prepared by George [5] in 36-percent yield by the reaction of *n*-propyl-magnesium halide and diisopropyl ketone. The residue (50 g) was a viscous material (boiling point >200° C, n_D^{20} 1.4590), and presumed to be a bimolecular product from condensation of diisopropyl ketone.

2. Diisopropyl Ketone and Isobutyl Chloride

A solution of 155 g of isobutyl chloride (1.68 moles) and 96 g of diisopropyl ketone (0.84 mole) was diluted with 200 ml of isooctane and added

¹ This reaction is visualized as starting with the formation of a sodium ketyl as the first step. This ketyl is colored an intense dark blue, blue-green, purple, or black, depending, it seems, on one or a combination of several factors. It has been noted that in several reactions tried, where reaction failed to occur, no characteristic color was observed. It is significant that this color appears first on the surface of the sodium. At this stage, if only a very small amount of ketone is present in relation to the amount of halide, the color often disappears as the ketone reacts.

TABLE I. Reaction products

Reaction	Product	Yield percentage of theory ¹	Boiling point	Refractive index, n_D^{20}	Density, D_4^{20}	Reference
Diisopropyl ketone and <i>n</i> -propyl chloride.	2-methyl-3-isopropyl-3-hexanol.....	32	^a 90 to 95/22.....	1.4438	0.8504
			^b 178 to 182/780.....			
Diisopropyl ketone and <i>n</i> -propyl chloride.	2-methyl-3-isopropyl-3-hexanol.....	32	^c 152/125.....	1.4444	.8554	[5]
			^d 182/780.....			
			^e 105 to 110/38.....	1.4430	0.8496
Diisopropyl ketone and isobutyl chloride.	2,5-dimethyl-3-isopropyl-3-hexanol.....	14	180 to 184/760.....	1.4483	.8470	[6]
			^f 140/125.....			
Diisopropyl ketone and isobutyl chloride.	2,5-dimethyl-3-isopropyl-3-hexanol.....	14	^g 160/760.....	1.4483	.8470	[6]
			^h 111 to 112/123.....	1.4397	0.8550
Pentamethylacetone and methyl bromide.	2,2,3,4-tetramethyl-3-pentanol.....	43	175/780.....	1.4405	.8565	[8]
			ⁱ 84/60.....			
Pentamethylacetone and methyl bromide.	2,2,3,4-tetramethyl-3-pentanol.....	29	^j 98 to 90/55.....	1.4410	0.8555
			^k 84/60.....	1.4405	.8565	[8]
Pentamethylacetone and ethyl bromide.	2,2,4-trimethyl-3-ethyl-3-pentanol.....	44	^l 93 to 97/35.....	1.4478	0.862
			^m 185 to 189/780.....			
Pentamethylacetone and ethyl bromide.	2,2,4-trimethyl-3-ethyl-3-pentanol.....	44	ⁿ 188 to 191/780.....	1.4481	.8628	[14]
			^o 105 to 107/35.....	1.4060
Pentamethylacetone and isopropyl chloride.	2,2,4-trimethyl-3-isopropyl-3-pentanol.....	42	^p 108 to 104/27.....	1.4559	[1]
			^q 117/22.....	1.4563	0.8505
Hexamethylacetone and <i>n</i> -propyl chloride.	2,2-dimethyl-3- <i>n</i> -butyl-3-hexanol ^a	88	117/22.....	1.4563	0.8505
			^r 103 to 107/30.....	1.4487	.8439
Pinacolone and <i>n</i> -butyl chloride.	2,2,3,3,6-pentamethyl-3-heptene-5-one.....	183/10.5.....	^s 1.4498	^t 0.8432
			^u	1.4498	0.8432
Hexamethylacetone and <i>s</i> -butyl chloride.	No product.....
Methyl <i>n</i> -butyrate and <i>t</i> -butyl chloride (small run).	2,2-dimethyl-3-hexanone ^a	9	^v 48 to 52/30.....	1.4115	[9]
			^w 148 to 152/760.....			
Methyl <i>n</i> -butyrate and <i>t</i> -butyl chloride (small run).	2,2-dimethyl-3-hexanone.....	6.5	^x 145 to 148/760.....	1.4065
			^y 76.2 to 77.7/86 to 89.....	1.4098
Methyl <i>n</i> -butyrate and <i>t</i> -butyl chloride (large run).	2,2-dimethyl-3-hexanone.....	6.5	^z 148 to 148/760.....	1.4095	[6,12,10]
			^{aa} 70 to 78/28 to 30.....	1.4262
Methyl <i>n</i> -butyrate and <i>t</i> -butyl chloride (large run).	2,2-dimethyl-3-hexanone.....	4.5	^{ab} 156/760.....	1.4262
			^{ac} 90 to 91/87 to 89.....			
Methyl <i>n</i> -butyrate and <i>t</i> -butyl chloride (large run).	2,2-dimethyl-3-hexanone.....	4.5	^{ad} 70 to 78/28 to 30.....	1.4280 (at 17° C).	[6,12,10]
			^{ae} 155 to 157/760.....			
Methyl isobutyrate and <i>t</i> -butyl chloride.	Hexamethylmethane.....	50	99 to 102.....	Solid
			^{af} 82/89.....			
Methyl isobutyrate and <i>t</i> -butyl chloride.	Pentamethylacetone.....	50	^{ag} 138/780.....	1.4059	[8]
			^{ah} 165/760.....	1.4260
Methyl isobutyrate and <i>t</i> -butyl chloride.	2,2,4-trimethyl-3-pentanol.....	7.3	^{ai} 83/82.....	1.4280
			^{aj} 102/760 ¹			
Methyl isobutyrate and <i>t</i> -butyl chloride.	2,2,4,4-tetramethyl-3-isopropyl-3-pentanol.....	19	^{ak} 113/37.....	1.4665	[2]
			^{al} 180 to 188/80.....	1.4289
Diethylcarbonate and <i>t</i> -butyl chloride.	Ethyl trimethylacetate.....	9	^{am} 118.....	1.3912	[6]
			^{an} 118.....	1.3906
Diethylcarbonate and <i>t</i> -butyl chloride.	Hexamethylacetone.....	9	^{ao} 151.....	1.4398
			^{ap} 155.....	1.4388
Diethylcarbonate and <i>t</i> -butyl chloride.	Pivaloin.....	Melting point, 78° C.....

See footnotes at end of table.

TABLE I. Reaction products—Continued

Reaction	Product	Yield percentage of theory ¹	Boiling point ° C/mm Hg	Refractive index, n_D^{20}	Density, D_4^{20}	Reference
Ethyl trimethylacetate and isopropyl chloride.	Pentamethylacetone.....	4	136 <i>136</i>	1.4059 <i>1.4059</i>		[8]
	2,2,4-trimethyl-3-pentanol.....	4	151 <i>151</i>	1.4287 <i>1.4287</i>		[9]
	2,2,4-trimethyl-3-isopropyl-3-pentanol.....	24	110/20 <i>105 to 104/27</i>	1.4576 <i>1.4576</i>		[1]
Panlaldehyde and <i>t</i> -butyl chloride.....	No product.....					
Panlaldehyde and <i>t</i> -butyl chloride.....	do.....					
Acetamide and <i>t</i> -butyl chloride.....	do.....					
Acetic anhydride and <i>t</i> -butyl chloride.....	do.....					

¹ Yields based upon the carbonyl or ester compound.

² The temperature conversion to 760 mm Hg was done by a method described by Lippencott [11].

³ The data in italics are from the literature.

⁴ A freezing point of -10.1°C for this compound was made by E. H. Rich.

⁵ The values for a known sample of 2,2,3,6-pentamethyl-3-heptene-5-one, as measured in this laboratory.

⁶ Obviously, a mixture of ketone and carbinol.

Constants from the carbinol previously prepared in this laboratory.

over a period of 3 hours to a suspension of 69 g of sodium sand (3 g atoms) dispersed in 1 liter of isooctane at 80°C . The reaction mixture turned purplish blue as soon as the addition was begun, and the temperature rose quickly to 98°C . Stirring was continued for 1 hour after the addition was completed. The reaction mixture was hydrolyzed after standing overnight, and the organic layer washed several times with water and dried. The solvent and unreacted isobutyl chloride were distilled off, and the residue was distilled under reduced pressure (Claisen flask) to give 20 g (14% yield) of 2,5-dimethyl-3-isopropyl-3-hexanol.

3. Pentamethylacetone and Methyl Bromide

A suspension of 69 g of sodium sand (3 g atoms) in a solution of 128 g of pentamethylacetone (1 mole) and 1 liter of isooctane was heated to 80°C . An excess of methyl bromide was added from a methyl bromide dispenser¹ through a tube extending below the surface of the liquid. The reaction began in the usual manner, with the temperature rising quickly to 98°C , and the appearance of the blue color. Stirring was continued for 1 hour after all of the methyl bromide

had been added, and the contents of the flask was allowed to stand overnight. The reaction mixture was worked up in the usual manner, and the isooctane and unreacted methyl bromide were removed by distillation. The residue, which was distilled at reduced pressure (Podbielniak still), yielded 62 g of 2,2,3,4-tetramethyl-3-pentanol, (0.43 mole, 43% yield).

4. Pentamethylacetone and Methyl Iodide

Pentamethylacetone (128 g, 1 mole) and 284 g methyl iodide (2 moles) were diluted with 200 ml of isooctane and added over a period of 2 hours to a suspension of 69 g of sodium sand (3 g atoms) in 300 ml of isooctane at 80°C . A rapid rise in temperature occurred, followed by the appearance of the characteristic blue color. After standing overnight the reaction mixture was hydrolyzed, and the organic layer washed several times with water and dried. The solvent and unreacted methyl iodide were removed by distillation, and the residue, which was distilled under reduced pressure, (Claisen flask) gave 41.5 g (29% yield) of 2,2,3,4-tetramethyl-3-pentanol.

5. Pentamethylacetone and Ethyl Bromide

A solution of 163 g of ethyl bromide (1.5 moles) and 64 g of pentamethylacetone (0.5 mole) was

¹ Arrow Products Co., Carlstadt, N. J.

added to 72 g of sodium sand (3.1 g atoms) suspended in 900 ml of isooctane. The reaction began in the usual manner with the appearance of the blue color and a rise in the temperature from 80° to 98° C. After standing overnight the reaction mixture was hydrolyzed, and the organic portion was washed and dried. Isooctane was removed by distillation, and the residue was distilled at a reduced pressure (Claisen flask) to give 35 g (44% yield) of 2,2,4-trimethyl-3-ethyl-3-pentanol. Nasarow [14] prepared the same carbinol in 7-percent yield by using pentamethylacetone and ethylmagnesium bromide.

A large run of this reaction has been carried out as a part of a synthesis. More precise data on the physical properties of the carbinol, its dehydration, and identification of 4,4-dimethyl-3-isopropyl-2-pentene as the dehydration product will be reported when further work, which is now in progress, has been completed.

6. Pentamethylacetone and Isopropyl Chloride

A solution of 64 g of pentamethylacetone (0.5 mole) and 98 g of isopropyl chloride (1.25 mole) was diluted with 250 ml of isooctane and added to a reaction flask containing 72 g (3.1 g atoms) of sodium sand in 300 ml of isooctane at 80° C. The usual procedure was carried out, and the organic layer was worked up in the usual manner. After removing the solvent, the residue was distilled at reduced pressure (Claisen flask) to give 36 g (42% yield) of 2,2,4-trimethyl-3-isopropyl-3-pentanol. This carbinol was prepared by Bartlett and Schneider [1] in 6-percent yield from the reaction of diisopropyl ketone with *t*-butyl chloride by means of sodium.

7. Hexamethylacetone and *n*-Propyl Chloride

A solution of 142 g of hexamethylacetone (1 mole) and 157 g of *n*-propyl chloride (2 moles) was diluted with 200 ml of isooctane. This solution was added to 69 g of sodium sand (3 g atoms) in 600 ml of isooctane at 80° C over a period of 2 hours. A sudden rise in temperature occurred and was followed by the appearance of the usual blue-black color in the reaction mixture. After standing overnight the contents of the flask was hydrolyzed, washed, and dried. Isooctane and unreacted *n*-propyl chloride were removed, and the residue was distilled at reduced pressure in a Claisen flask. There was obtained a 38-percent

yield of 2,2-dimethyl-3-*t*-butyl-3-hexanol. Two analyses⁵ on this carbinol gave (1) C, 77.6 percent; H, 14.3 percent and (2) C, 77.6 percent; H, 14.7 percent. The calculated values are: C, 77.42 percent; H, 13.98 percent. The molecular refractions were determined: MR (observed), 58.97, MR (calculated),⁶ 58.86.

It is presumed that the carbon and hydrogen analyses are high due to the presence of some olefinic material formed by the dehydration of the carbinol during distillation.

8. Hexamethylacetone and *t*-Butyl Chloride

A solution of 71 g of hexamethylacetone (0.5 mole) and 138 g of *t*-butyl-chloride (1.5 moles) was added to 69 g of sodium sand (3 g atoms). The reaction began at 78° to 80° C, but it did not proceed well. Considerable gas, probably isobutylene, was given off during the reaction. The ketone-halide solution was added to the reaction mixture over a period of 2 hours. The contents of the flask was worked up in the usual manner, and the solvent was removed at atmospheric pressure. The residue (85 ml) was distilled from a Claisen flask at reduced pressure to give: 25 ml of residual isooctane, 36 g of recovered hexamethylacetone, boiling point 65° to 72° C at 32 mm Hg, n_D^{20} = 1.4133 to 1.4195, and 3 ml of heavy viscous material of unknown composition. An attempt to crystallize tri-*t*-butyl carbinol [1] from a pentane solution of the residue by using first an ice-salt mixture and secondly a dry ice and acetone mixture, failed to yield any solid material.

9. Pinacolone and *n*-Butyl Chloride

A solution of 100 g of pinacolone (1 mole) and 185 g of *n*-butyl chloride (2 moles) in 700 ml of isooctane was added to 69 g of sodium sand (3 g atoms) in 300 ml of isooctane at 80° C over a period of 2 hours. The reaction that ensued was similar to those already described. After standing overnight, the reaction mixture was worked up in the previously described manner. The isooctane and unreacted *n*-butyl chloride were removed at atmospheric pressure, and the residue was distilled at reduced pressure in a Claisen flask. There was obtained 43 ml of material having these properties: boiling point, 103° to

⁵ We are indebted to the Analytical Chemistry and Standard Samples Sections, NBS, for the carbon and hydrogen determination.

⁶ Lorentz-Lorentz equation.

107° C at 30 mm Hg, $n_D^{20}=1.4487$, $D^{20}=0.8439$. The reported constants for the expected product, 2,2,3-trimethyl-3-heptanol, are boiling point, 84° to 87° C at 13 mm Hg, n_D^{20} 1.4409, D^{20} 0.8487 [3]. Since the constants of the fraction obtained do not agree with the reported values, it was presumed that pinacolone had dimerized in the presence of sodium to yield, after dehydration, 2,2,3,6,6-pentamethyl-3-heptene-5-one, boiling point, 83° C at 10.5 mm Hg, $n_D^{21}=1.4498$, $D^{21}=0.8432$. A carbon and hydrogen determination gave C, 79.1 percent; H, 12.6 percent. The calculated values for the dimer are: C, 79.12 percent; H, 12.08 percent, and for 2,2,3-trimethyl-3-heptanol, C, 76 percent; H, 13.92 percent. The formation of the dimer was expected rather than the addition reaction, because pinacolone, being a methyl ketone, easily aldolizes in the presence of sodium.

A 2,4-dinitrophenylhydrazone was prepared from some of this material, which melted 146.5° to 147° C. The same derivative was prepared from an authentic sample of 2,2,3,6,6-pentamethyl-3-heptene-5-one and melted at 147° to 147.5° C. A mixture of the two derivatives melted at 146.0° to 146.5° C. Hence the main product from this reaction was the dehydrated bimolecular product from pinacolone.

10. Methyl *n*-Butyrate and *t*-Butyl Chloride

A solution of 102 g of methyl *n*-butyrate (1 mole) and 368 g of *t*-butyl chloride (4 moles) was added over a period of 3 hours to 46 g (2 g atoms) of sodium sand in 200 ml of isooctane at 80° C. After standing overnight the mixture was hydrolyzed, and the organic layer was washed several times with water and dried over anhydrous sodium carbonate. The isooctane was removed, and the residue was distilled at reduced pressure in a Claisen flask. There was obtained thereby 11 g (9% yield) of 2,2-dimethyl-3-hexanone.

Since more information was desired on this reaction than was afforded by the small-scale run, a large run was made. A solution of 17.6 kg of *t*-butyl chloride (190 moles) and 7.54 kg of ethyl *n*-butyrate (65 moles) was added over a period of 7 hours to 6.9 kg of sodium sand (300 g atoms) in 10 gal of isooctane at a temperature of 55° to 60° C under a nitrogen atmosphere. The reaction mixture turned greenish-black during the course of the addition and evolved large quantities of gas,

probably isobutylene. The reaction mixture was stirred overnight at a temperature of 56° C.

Five gallons of 95-percent ethanol was then added over a half-hour period, and the mixture was kept at reflux temperature all day in order to remove as much as possible of the excess sodium. This was followed by 10 gal of water, which was added cautiously to the cooled reaction mixture under a nitrogen atmosphere during 6 hours. After the last traces of sodium had disappeared, the mixture was washed three times with a large amount of water. About 8 to 10 gal of organic material was obtained, which was dried over calcium chloride prior to distillation.

The isooctane was removed at atmospheric pressure, and the residue, 2,500 ml, was submitted to vacuum distillation (Podbielniak). There were obtained 552 g of 2,2-dimethyl-3-hexanone (4.3 moles, 6.6%) and 408 g of 2,2-dimethyl-3-hexanol (3.1 moles, 4.8%) [10]. There were no fractions that could be identified as the desired product, 2,2-dimethyl-3-*t*-butyl-3-hexanol.

The poor recovery of starting materials is indicative of the fact that little coupling occurred; the boiling points of the coupling products are all well above that of isooctane, and they would certainly have appeared in the residue left upon stripping the isooctane. Hence their absence can only be accounted for as a consequence of side reactions. The *t*-butyl chloride was lost partly as isobutylene during the course of the reaction, and the remainder distilled with the isooctane during the stripping operation. The ester must have been lost as a consequence of alkaline hydrolysis when the reaction mixture was worked up, the hydrolysis products being discarded with the raffinates.

11. Methyl Isobutyrate and *t*-Butyl Chloride

In a 50-gal stainless steel reactor, 6.9 kg of sodium sand (300 g atoms) was prepared in 10 gal of isooctane. To this was added 17.6 kg of *t*-butyl chloride (191 moles), and the mixture was heated to 45° C, after which 6.4 kg of methyl isobutyrate (62.8 moles) was added over a period of 6 hours. The reaction mixture was stirred overnight and hydrolyzed. The organic layer was separated, washed, and dried. Isooctane and unreacted *t*-butyl chloride were distilled at atmospheric pressure. The following compounds were isolated from the residue by distillation on a

Podbielniak column: 59 g of hexamethylethane, 4,036 g of pentamethylacetone (50% yield), 599 g of 2,4,4-trimethyl-3-pentanol, and 2,229 g of 2,2,4,4-tetramethyl-3-isopropyl-3-pentanol (19% yield). The latter carbinol was prepared in 15-percent yield by Bartlett and Schneider [1] by the reaction of pentamethylacetone, *t*-butyl chloride, and sodium.

12. Diethylcarbonate and *t*-Butyl Chloride

Ethyl trimethylacetate (9% yield) and hexamethylacetone (9% yield) were obtained from the reaction of 277 g of *t*-butyl chloride (3 moles) and 236 g of diethylcarbonate (2 moles) in the presence of 92 g of sodium sand (4 g atoms) using petroleum ether as the solvent. There was considerable difficulty in starting the reaction, but once started the reaction became very vigorous. The sodium turned black, then bronze, and the reaction mixture thickened considerably so that additional solvent was required to permit stirring. After hydrolyzing the reaction mixture, the organic layer was worked up in the usual manner. After removing the petroleum ether, the residue was distilled in a Podbielniak column to give 0.185 mole of ethyl trimethylacetate, and 0.176 mole of hexamethylacetone.

A second run was made by using the same quantity of materials, but isooctane was used as the solvent. From the residue 25 g of solid (melting point, 78° C), which was identified as pivaloin (melting point, 80° C), and a very high boiling material were obtained. There were no fractions isolated having the properties of either ethyl-trimethylacetate or hexamethylacetone.

13. Ethyl Trimethylacetate and Isopropyl Chloride

In a 50-gal stainless steel reactor, 15.4 lb of sodium sand (300 g atoms) was prepared in 10 gal of isooctane. To this there was added, over a period of 6 hours, a solution of 180 moles of isopropyl chloride and 60 moles of ethyl trimethylacetate. The reaction mixture exhibited the characteristic appearance of a blue color accompanied by a rise in the temperature to 54° C. The reaction was kept at 50° to 70° C by means of warm water in the kettle jacket. After standing overnight, excess sodium was destroyed by 5 gal of 190-proof ethyl alcohol, followed by ample quantities of water for hydrolysis. The organic

layer was separated, washed several times with water, and dried over anhydrous calcium chloride. Isooctane was distilled off, and the residue was distilled at reduced pressure in a Podbielniak column. The following products were isolated and identified from their physical constants: 2.5 moles (4% yield) pentamethylacetone, 2.73 moles (4% yield) 2,2,4-trimethyl-3-pentanol, and 32.6 moles (54% yield) 2,2,4-trimethyl-3-isopropyl-3-pentanol. Distillation data are shown in figure 1.

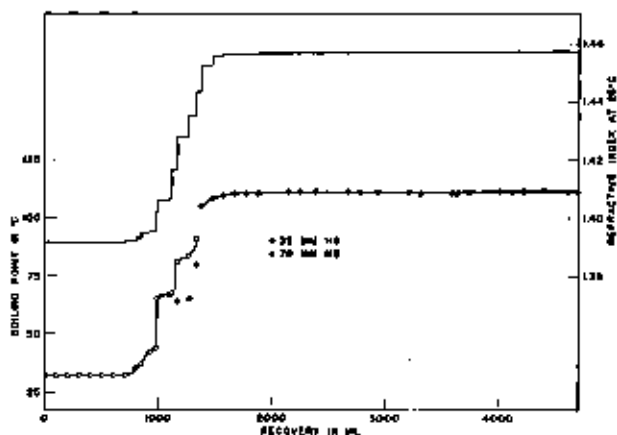


FIGURE 1. Results of distillation of the products from interaction of ethyl trimethylacetate and isopropyl chloride.

The small plateau in the boiling point curve, beginning at about 1,000 ml. distilled, represents pentamethylacetone; the next small plateau represents 2,2,4-trimethyl-3-pentanol. The long plateau represents 2,2,4-trimethyl-3-isopropyl-3-pentanol. These curves are similar to those obtained in the other experiments.

14. Paraldehyde, Paraformaldehyde, and *t*-Butyl Chloride

An attempt was made to extend the Bartlett-Schneider reaction to include the reaction of an aldehyde with an alkyl halide. Because monomeric aldehydes aldolize under alkaline conditions, it was decided to try paraldehyde and paraformaldehyde.

A mixture of 132 g of paraldehyde (0.5 mole) and 138 g of *t*-butyl chloride (1.5 mole) was added in small portions to 39 g of sodium sand in 400 ml of isooctane. The temperature of the reaction mixture was maintained at 80° C until all of the mixture had been added. There was no evidence of reaction at this temperature; consequently the mixture was heated to 90° C for 1 hour. After standing overnight, no reaction had taken place. The experiment was discontinued.

A similar attempt was made with paraformaldehyde and *t*-butyl chloride as reactants, but no reaction was observed.

15. Acetamide and *t*-Butyl Chloride

Small portions of 59 g of acetamide (1 mole) and 277 g of *t*-butyl chloride (3 moles) were added to a reaction vessel containing 69 g of sodium sand (3 g atoms) in 1 liter of isooctane. The reaction was very sluggish with evolution of a gas (probably isobutylene), and the temperature did not rise above the initial temperature of 80° C. After hydrolysis with ethyl alcohol and water, a strong odor of ammonia developed indicating hydrolysis of the acetamide to acetic acid. Fractionation gave a cut (boiling point 78° C $n_D^{20}=1.3862$) identified as ethyl acetate [9]. This is not surprising, since acetic acid and ethyl alcohol could esterify under conditions of the experiment. No other reaction product was obtained.

16. Acetic Anhydride and *t*-Butyl Chloride

In a large Erlenmeyer flask, 94 ml of acetic anhydride was added to 23 g of sodium sand in 100 ml of isooctane. The mixture was heated to 75° C, and 217 ml of *t*-butyl chloride was added in small amounts. From time to time the flask was heated on a hot plate. There were no visible signs of reaction, nor did any reaction occur when the contents of the flask was heated to the boiling point of isooctane (boiling point, 99° C). The addition of *t*-butyl chloride was discontinued, and the reaction mixture was allowed to stand overnight. No change had taken place during the night, and the experiment was therefore discontinued.

III. Summary

In the study of these condensation reactions, seventeen runs were made, of which nine were of

the aliphatic ketone-aliphatic halide type, four were of the aliphatic ester-aliphatic halide type, and four were a miscellaneous type. In those reactions that produced desirable compounds, the yields were as good or better than where the Grignard reagent was employed.

This reaction is apparently, applicable to compounds whose carbonyl group is sufficiently removed from influences that cause enolization, aldolization, or hydrolysis. Table I is a résumé of results obtained in this work.

IV. References

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