NIST Technical Note 2082

The Structure of Medium-Scale Pool Fires

Ryan Falkenstein-Smith Kunhyuk Sung Jian Chen Kimberly Harris Anthony Hamins

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Ryan Falkenstein-Smith Kunhyuk Sung Jian Chen Anthony Hamins Fire Research Division Engineering Laboratory

Kimberly Harris Chemical Science Division Material Measurement Laboratory

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Abstract

This report documents a series of time-averaged gas species measurements made along the centerline of methanol, ethanol, acetone, and methane pool fires steadily burning in a quiescent environment. All gas species measurements are obtained using a Gas Chromatograph/ Mass Spectrometer System (GC/MS) for gas samples extracted at various heights above the fire and repeated at least twice at each location. Gas species volume fractions are determined via the GC/MS using predetermined calibration factors. Soot mass fractions are simultaneously measured during the gas sampling process. The gas species volume and soot mass fractions are compared at different heights within the fire and across a variety of different fuels. Other fire parameters are measured as well, including time-averaged temperature measurements and mass burning rates.

Keywords

Acetone, Ethanol, Gas pool fires, Gas species measurements, Methanol, Methane, Liquid pool fires

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Nomenclature

Nomenclature			Amount of moles injected into
A	Peak area		Total amount of moles
A _{corr}	Corrected peak area	n _{tot}	Neces la seconde a
С	Total number of carbon atoms	Nu	Nusselt number
$C_{\rm vap}$	Volume fraction of vapors	P	Pressure
$c_{p_{air}}$	Specific heat of ambient air	P_{∞}	Ambient Pressure
сb	Specific heat of thermocouple bead	P _{vap}	Vapor Pressure
D	Diameter of Pool Fire	φ	Local equivalence Ratio
D		Pr	Prandtl number
D_{b}	Diameter of thermocouple bead	Ż	Heat release rate
ε	Thermocouple emissivity	R	Universal gas constant
η	Reciprocal of the local fuel equiva-	Re	Reynolds number
		$ ho_{_{ m air}}$	Density of air
8	Acceleration of gravity	$ ho_{ m b}$	Density of thermocouple bead
ΔH_c	Heat of combustion	Øα	Density of gas
Η	Total number of hydrogen atoms	$\rho_{\infty_{aix}}$	Density of ambient air
h	Convective heat transfer coefficient	S _i	Standard deviation of component i
kg	Thermal conductivity of the gas	σ	Stefan-Boltzmann constant
L_{f}	Mean flame height	Т	Temperature
ṁ	Mass burning rate	Tb	Measured thermocouple bead tem-
m _s	Mass of soot		perature
m _t	Total mass of gas injected into	Tg	Gas temperature
	GC/MS	T_{∞}	Ambient Temperature
m _{tot}	Total mass of gas sampled	t	time
$\mu_{ m air}$	Dynamic viscosity of air	$U_{ m g}$	Velocity of gas
n _i	Amount of moles of species i	<i>u</i> _i	Uncertainty of component i

	Volumetric flow rate	$ar{Y}_{ m F}$	Mass fraction of fuel molecule
$V_{\rm s}$	Volume of sampled gas injected into GC/MS	\bar{Y}_i	Mean mass fraction
W _F	Molecular weight of fuel molecule	Уi	Numbers of hydrogen atoms in the molecule
W _i	Molecular Weight of species i	Ys	Mass fraction of soot
\bar{X}_i	Mean volume fraction	Ζ	Mixture Fraction
X	Number of carbon atoms in the fuel molecule	Z.	Vertical spatial coordinate
X _i	Numbers of carbon atoms in the molecule	z^*	Normalized vertical spatial coordinate

1. Introduction

Computational fluid dynamics (CFD) models are an important component of performancebased design in fire protection engineering. A requirement of their acceptance in the design process is that these models be verified and validated, the latter of which involves comparison with experimental measurements. The primary objective of this report is to provide data for use in fire model validation.

A pool fire is a fundamental focus of study in fire science. The fuel surface is isothermal, flat and horizontal, providing a well-defined boundary condition for modeling. Fuel and product species concentrations and temperatures have a significant influence on the heat feedback to the fuel surface, which directly affects the burning rate. A zone of particular interest is the fuel rich-core between the flame and the pool surface, where gas species can absorb energy that would otherwise have been transferred to the fuel surface. Few studies in the literature have reported local chemical species measurements within the flame envelop.

The purpose of this study is to characterize the spatial distribution of the principal chemical species in moderate-scale liquid pool fires steadily burning in a well-ventilated, quiescent environment. Here, methanol, ethanol, acetone, and methane are the fuels of interest. Contrary to ethanol and acetone, fires established using methanol are unusual as no carbonaceous soot is present or emitted. These particular fuels are selected for research since the measurements complement results from previous studies, including analyses of the mass burning rate, the temperature and velocity fields, radiative emission, flame height, and pulsation frequency [1, 2].

2. Description of Experiments

Descriptions of previous pool fire experiments are found in Refs. [2–6]. All experiments are conducted under a canopy hood surrounded by a cubic enclosure, 2.5 m on a side, made of a double layer wire-mesh screen (5 mesh/cm) to reduce the impact of room ventilation. All measurements are made once the mass burning rate reaches a steady-state, achieved approximately 10 min and 2 min after ignition for the liquid and methane fuels, respectively.

2.1 Liquid Pool Burner Setup

The circular, stainless-steel burn pan, made from cold rolled steel, has an outer diameter of 30 cm, a depth of 15 cm, and a wall thickness of 1.6 mm. The lip of the burner is positioned 30 cm above the floor. As shown in Fig. 1, the burner is placed within an overflow basin, which extends 3 cm beyond the burner wall. The bottom of the burner is maintained at a constant temperature by flowing water (approximately $20 \degree C \pm 5 \degree C$) through the basin. A fuel level indicator is positioned near the center of the burner to maintain a consistent surface level.



Fig. 1. The 30 cm burner with fuel level indicator, overflow section, and quenching probe

Fuel to the burner is gravity fed from a reservoir positioned on a mass load cell located outside the enclosure and monitored by a data acquisition system (DAQ). As shown in Fig. 2, the fuel level is monitored via a fuel flow operator. The operator is able to observe a close up of a slightly discernible dimple (approximately 2 mm) made from the fuel level

indicator on the fuel surface using a live video feed. The fuel level is controlled by manually adjusting the fuel flow using a needle valve. The fuel surface is maintained 10 mm below the burner rim to match previous experimental conditions [1, 2, 7, 8].

2.2 Gas Pool Burner Setup

Gas fuels are burned using a 38 cm diameter burner. Fuel to the gas burner is controlled via a mass flow controller located outside of the enclosure. Similar to the liquid pool burner, the gas burner is maintained at a constant temperature by circulating cooled water (approximately $20 \degree C \pm 5 \degree C$) through the basin.

2.3 Measuring Flame Characteristics

The heat release rate, \dot{Q} , is the product of the time-averaged mass burning rate, \dot{m} , and the heat of combustion, ΔH_c :

$$\dot{Q} = \dot{m} \,\Delta H_c \tag{1}$$

The heat of the combustion is provided by Ref. [9] and reported in Table 1.

The mean flame height, L_f , is estimated from 3600 frames of high-resolution video of the experiments using MATLAB's Image Processing Toolbox¹. Imported color images are decomposed into binary (i.e., black and white) images using a pre-set threshold level. The flame height for a single frame is defined as the distance between the pool surface and flame tip. All measurements are repeated, then averaged to provide the mean flame height. A description of the uncertainty analysis for the mean flame height is described in Appendix A.3.

2.4 Centerline Temperature Measurements

Time-averaged temperature measurements are made along the vertical centerline of the fire plumes using S-type (Pt 10% Rh/Pt), bare-wire, fine diameter thermocouples (OMEGA P10R-001) with wire diameters of approximately 13 μ m and 25 μ m for the liquid and methane pool fires, respectively, and bead diameter approximately three times greater. Temperature measurements are sampled at 250 Hz for 2 min, or approximately 300 pulsing cycles [10]. The uncertainty of the temperature radiative measurements are discussed in Appendix A.4.

The measured thermocouple temperatures are corrected for heat losses and thermal inertia using the formula described by Shaddix [11]:

$$T_{\rm g}(t) = T_{\rm b}(t) + \tau \frac{\mathrm{d}T_{\rm b}}{\mathrm{d}t} + \frac{\varepsilon\sigma}{h} \left(T_{\rm b}(t)^4 - T_{\infty}^4\right) \tag{2}$$

¹ Certain commercial products are identified in this report to specify adequately the equipment used. Such identification does not imply a recommendation by the National Institute of Standards and Technology, nor does it imply that this equipment is the best available for the purpose.



Fig. 2. Photo of fuel flow operator monitoring the fuel level via live video feed (top) and a magnified image of the live video feed used to maintain a consistent fuel level relative to the fuel level indicator (bottom)

where T_g is the "true" gas temperature, T_b is the measured bead temperature, T_{∞} is the ambient temperature, $\sigma = 5.67 \times 10^{-11} \text{ kW}/(\text{m}^2\text{K}^4)$ is the Stefan-Boltzmann constant,

 ε is the thermocouple emissivity, and *h* is the convective heat transfer coefficient. The temperature-dependent emissivity of the platinum is taken from Ref. [11] and shown below:

$$\varepsilon = -0.1 + 3.24 \times 10^{-4} T - 1.25 \times 10^{-7} T^2 + 2.18 \times 10^{-11} T^3$$
(3)

h and τ are defined:

$$h = \frac{\operatorname{Nu}k_{g}}{D_{b}} \quad ; \quad \tau = \frac{\rho_{b}c_{b}D_{b}^{2}}{6\operatorname{Nu}k_{g}} \tag{4}$$

where k_g is the thermal conductivity of the gas, ρ_b , c_b , and D_b are the density, specific heat, and diameter of the bead, respectively. The thermal conductivity of the gas is determined from datasets provided in Ref. [12]. The density of the thermocouple bead is assumed constant at 21.45 g/cm [13]. The specific heat of the bead is calculated from the formulae provided in Ref. [14]. The Nusselt number, Nu, is calculated using the Ranz-Marshall correlation [11]:

Nu = 2 + 0.6 Re^{1/2} Pr^{1/3} ; Re =
$$\frac{\rho_{air} U_g D_b}{\mu_{air}}$$
 ; Pr = 0.7 (5)

The temperature-dependent gas properties for Reynolds number, Re, and Prandtl number, Pr, are taken as those of air from Ref. [15]. The gas velocity is assumed to be equal to 2 m/s. The corrected temperature is relatively insensitive to gas velocities between 1 m/s and 3 m/s, consistent with the results of Shaddix [11].

The uncertainty of Eq. (2) is discussed in detail in Ref. [16]. This correction formula is included for the convenience of mathematical modeling.

2.5 Measuring the Volume Fraction of Gas Species via GC/MS

Gas species measurements are made using an Agilent 5977E Series GC/MS fitted with a thermal conductivity detector (TCD). The GC/MS is able to quantify a variety of stable reactant, intermediate, and product species collected from the fire plume. Figure 3 displays the flow diagram for gas sampling. The GC/MS is equipped with a 2 mL sample loop maintained at approximately 200 °C. The gases are extracted by a vacuum pump located downstream of the GC/MS. Gas samples are collected using a quenching probe, which is composed of two concentric, stainless-steel tubes with outer annular coolant flow and inner, extracted sample flow. The inner and outer tube diameters are 8 mm and 16 mm, respectively. Water at approximately 90 °C flows through the sampling probe for the duration of the experiment. The remainder of the sampling line leading into the GC/MS is heated with electrical, heating tape to approximately 140 °C to prevent condensation of water and liquid fuels within the line.

Depending on the probe location within the fire, the sampling period varies from 12 min to 25 min, ensuring that the gases are completely swept through the sample loop. The flow is controlled using a mass flow controller (Alicat Scientific MC-Series) located in front of the vacuum pump within the sampling line. During the gas sampling procedure, the volumetric flow is approximately 200 mL/min and recorded at 2 Hz.







After the gas sampling period, two quarter-turn valves located on opposite ends of the sample loop are closed, at which point pressure measurements, obtained from a digital pressure gauge (OMEGA DPG409-030DWU), and temperature measurements, acquired by a K-type thermocouple located at the GC/MS sample loop injection port, are collected at 2 Hz for 50 s. The sampled gas is then fed onto a column located within the GC oven. Chromatographic separation of species is achieved using a Select for Permanent Gases-Dual Column (CP7430) comprised of mole-sieve and Porapak Q columns working in parallel and using a helium carrier gas. The flow path is divided past the column using a splitter, wherein the carrier gas flow leading into the TCD and MS is approximately 3 mL/min and 1 mL/min, respectively. When developing the GC/MS method, the portioned flows feeding into the TCD and MS are calculated such that their respective chromatograms are synchronized.

For the ethanol and acetone pool fires, the sample analysis time is approximately 60 min, during which time the GC oven temperature is maintained at 30 °C for 10 min, then ramped at 8 °C/min for 34 min until reaching a temperature of 300 °C which is maintined for the remainder of the analysis. For methanol pool fires, GC oven has same setpoint temperature hold time and ramp rate, but only ramps to 192 °C to reduce overall analysis time. During the sample analysis, the TCD is maintained at 300 °C with a makeup and reference flow of 12 mL/min and 27 mL/min, respectively. Additionally, the MS source and quad temperatures are 250 °C and 200 °C, respectively, for the duration of the sample analysis.

Gas species are identified using the total ion chromatogram (TIC) provided from the MS and quantified from the TCD chromatogram. A typical TCD chromatogram is shown in Fig. 4. The area associated with each peak is proportional to the number of moles of a particular species. Peaks are manually integrated using the Agilent software. Integration bounds of each peak are determined using the MS. In instances where peaks overlap, it is necessary to implement a tangential "skim" to improve integration resolution of an overlapping region, as shown in Fig. 4.

The area associated with a peak, A, is adjusted to account for the variation in the measured pressure, P, and temperature, T, of the sample:

$$A_{\rm corr} = A \left(\frac{P T_{\infty}}{P_{\infty} T}\right) \tag{6}$$

The corrected area is converted into number of moles using a linear calibration curve determined using gas standards listed in Appendix C.2. The calibration functions of liquidvapors (e.g., water, methanol, ethanol, acetone) are obtained from a bubbler apparatus described in Appendix C.3.

Once the number of moles for all detected species is determined, the volume fraction of each species can be calculated. The mean volume fraction of a given species, \bar{X}_i , is the ratio of the number of moles of a given gas species, n_i , and the total number of moles identified by the GC/MS, n_{tot} , averaged over the repeated GC/MS gas injections. The total number of moles is determined from the summation of moles for each species quantified from the TCD chromatogram. In the case of water, the relative humidity measured within the room



Fig. 4. TCD chromatogram of a pool fire gas sample and a magnified image of overlapping peaks with an example of tangential skim integration technique used to determine peak area

accounted for when calculating its volume fraction.

$$\bar{X}_i = \frac{n_i}{n_{\text{tot}}} \tag{7}$$

The mean mass fraction, \bar{Y}_i , of a given species *i* is calculated from the measured volume fraction, \bar{X}_i , using the following expression:

$$\bar{Y}_i = \frac{\bar{X}_i W_i}{\sum \bar{X}_i W_i} \tag{8}$$

where W_i is the molecular weight of a given species.

All measurements using the GC/MS are repeated at least twice at each location along the centerline of the pool fire. Gas species concentration measurements made at the same location are averaged. The variance in the volume fraction is a function of position and species. The uncertainty of the species measurements and the calibration procedure is discussed in Appendices B and C, respectively.

2.6 Determining Soot Mass Fraction

Soot mass fraction, Y_s , is measured using a well established gravimetric technique [17]. Soot is filtered out of the gas stream using a stainless steel particulate filter holder (PALL 2220). Before an experiment, a desiccated 47 mm polytetrafluoroethylene (PTFE) filter is weighed and placed into its holder. The filter holder is positioned within the gas sampling line behind the quenching probe and heated with tape to approximately 140 °C to prevent condensation of water and liquid fuels on the filter. After sampling, the filter is removed and dried in a desiccator. After drying for 48 h, the filter's final weight is measured. Approximately 1 mg of soot is collected during the sampling period, which varies from 12 min to 25 min depending on the sampling location. The mass of the PTFE filter and cleaning patches are measured three times before and after each test².

The soot mass fraction, Y_s , is computed from the mass of the soot collected from the PTFE filter and gun cleaning patches, m_s , the ratio of the mass flow controller's temperature reading, T_{∞} , to the effective temperature of the gas calculated from Eq. (2), T_g , the total mass of gas sampled, m_{tot} , based on the mass flow controller readings:

$$Y_s = \frac{m_{\rm s}}{m_{\rm tot}} \frac{T_{\infty}}{T_{\rm g}} \tag{9}$$

The total mass of gas sampled is the product of the average volumetric flow rate measured by the mass flow controller, \dot{V} , the density of the sample gas injected into the GC/MS, ρ_g , and the gas sampling time, Δt .

$$m_t = \dot{V} \rho_g \Delta t \tag{10}$$

In Eq. (10), the density of the sample gas is determined from the total mass detected from the TCD and MS, m_{tot} , for the injected sample volume, V_{s} .

$$\rho_{\rm g} = \frac{m_{\rm tot}}{V_{\rm s}} \tag{11}$$

A description of the soot mass fraction uncertainty is provided in Appendix D.

²After some experiments, soot deposits are observed on the inner walls of the quenching probe. As shown in Fig. 5, dedicated gun cleaning patches (Hoppe's 9 1203S) are used to clean the inside of the quenching probe with no cleaning solvent. At least two patches are used to collect soot on the inside of the probe. A petri dish is placed below one end of the probe to catch dislodged soot and patches. Soot collection on the inside of the probe concludes once an applied patch is observed to have no soot. Patches are weighed before and 48 h after cleaning the inside of the probe.



Fig. 5. Process of collecting soot from the internal walls of the quenching probe using gun cleaning patches

Table 1. List of measurements and thermochemical properties of fuels burning in a well-ventilated round 30 cm diameter pool fire burning in a quiescent environment. The uncertainty of the mass burning rate is discussed in Appendix A.1. The uncertainty of the heat release rate and \dot{Q}^* is discussed in Appendix A.2.

Parameter (units)	Methanol	Ethanol	Acetone	Methane
Mass Burning Flux (g/m ² s)	12.4 ± 1.1	13.9 ± 0.8	17.6 ± 2.7	6.4 ± 0.1
Heat Release Rate (kW)	17.4 ± 1.4	26.3 ± 1.5	35.5 ± 5.4	34.5 ± 0.5
\dot{Q}^*	0.32 ± 0.02	0.48 ± 0.01	0.64 ± 0.05	0.63 ± 0.01
<i>D</i> * (m)	0.190 ± 0.006	0.224 ± 0.005	0.252 ± 0.015	0.249 ± 0.003
Mean Flame Height (cm)	36.4 ± 16.0	61.1 ± 28.2	91.5 ± 34.6	64.0 ± 31.0
Heat of Combustion (kJ/g) [19]	19.94	26.81	28.56	50.03
Carbon/Hydrogen Ratio	1/4	1/3	1/2	1/4

3. Results

This section presents the flame height, temperature, gas species, and soot measurements made at incremental heights along the centerline of 30 cm diameter methanol, ethanol, and acetone pool fires and a 38 cm diameter methane pool fire.

3.1 Flame Observations

Figure 6 displays a series of snapshots depicting a single pulsation cycle of the methanol, ethanol, and acetone fires. The pulsation frequency is approximately 3 Hz. The methanol fire is purely blue, whereas the ethanol, acetone, and methane fires are more luminous and yellow. The measured time-averaged burning rates and calculated heat release rates are listed in Table 1. The heat release rates are calculated from Eq. (1). The methanol fire has the lowest average flame height, followed by the ethanol, methane, and acetone. The measured mean flame heights match Heskestad's correlation [18] to within measurement uncertainty. The measured flame heights are also within the uncertainty bounds of measurements made by Kim et al. [7].

$$\frac{L_{\rm f}}{D} = 3.7 \, (\dot{Q}^*)^{2/5} - 1.02 \quad ; \quad \dot{Q}^* = \frac{\dot{Q}}{c_p \rho_{\infty} T_{\infty} \sqrt{g} D^{5/2}} \tag{12}$$

Here, *D* is the diameter of the pool fire (30 cm), *g* is the acceleration of gravity, and c_p and ρ_{∞} are the specific heat and the density of air at room temperature, T_{∞} . The measured flame heights are also within the uncertainty bounds of measurements made by Kim et al. [7].



Fig. 6. Flame structures of methanol, ethanol, acetone, and methane pool fires during their pulsing cycles

3.2 Comparison of Pool Fires from Different Fuels

Figure 7 displays the time-averaged, corrected gas temperatures as a function of the normalized vertical spatial coordinate, z^* :

$$z^* = \frac{z}{D^*} \quad ; \quad D^* = \left(\frac{\dot{Q}}{c_p \rho_{\infty} T_{\infty} \sqrt{g}}\right)^{\frac{2}{5}} \tag{13}$$

Here, z is the vertical spatial coordinate, \dot{Q} is the heat release rate, g is the acceleration of gravity, and c_p and ρ_{∞} are the specific heat and the density of air at room temperature, T_{∞} . The maximum mean temperature for each fuel peaks at approximately $z^* = 0.2$. The maximum mean temperature for each fuel peaks is close to their respective stoichiometric mixture fraction values. Methane has the highest peak mean temperature of 1350 K with methanol, ethanol, and acetone exhibiting maximum mean temperatures of 1316 K, 1281 K, and 1190 K, respectively. The methanol temperature profile is in agreement with previous works [10]. The thermal inertia correction applied to the thermocouple temperature has less than a 5 K influence on the mean but significantly alters the RMS.



Fig. 7. Mean and root mean square (RMS) centerline temperature profiles of methanol, ethanol, acetone, and methane pool fires during their pulsing cycles

Figure 8 displays the mean volume fraction of the major species, \bar{X}_i , as a function of z^* for the methanol, ethanol, acetone, and methane fires. Plots for individual species, including uncertainties, are displayed in Appendix E. Major species detected in the TCD and MS include combustion reactants (fuels and oxygen, O₂), combustion products such as water, H₂O, and carbon dioxide, CO₂, combustion intermediates such as carbon monoxide,

CO, hydrogen, H_2 , and inert gases such as nitrogen, N_2 , and argon, Ar. Methane is detected and quantified in all fires. In the case of the ethanol and acetone fires, soot, benzene, acetylene, ethylene, and ethane are also detected and quantified. Trace amounts of other species are also detected including propene, acetaldehyde, and ethyl acetate, consistent with previous literature [20, 21].



Fig. 8. Centerline volume fraction and soot mass fraction profiles of methanol (\blacklozenge), ethanol (\bigcirc), acetone (\blacksquare), and methane (\bigtriangleup) pool fires.

4. Verifying Gas Species Measurements

This section presents several different techniques of verifying the accuracy of the species concentration measurements.

4.1 Counting Moles

As a way to verify the accuracy of the experimental method, specifically the calibration procedure and curve fit, the total moles, n_{tot} , identified by the TCD and MS is compared to the total moles injected into the GC/MS system, n_{inj} , which is calculated from the ideal gas law:

$$n_{\rm inj} = \frac{PV_{\rm s}}{RT} \tag{14}$$

Here, R = 8.314 J/(mol·K) is the universal gas constant, $V_s = 2 \times 10^{-6}$ m³ is the injected sample volume, and *P* (Pa) and *T* (K) are the mean pressure and temperature, respectively, collected before injecting the gas sample into the GC/MS as described in Section 2.5. The combined uncertainty of the total moles injected is detailed in Section C.1.



Fig. 9. Ratio of moles identified to moles injected, with uncertainty, as a function of z^* . The uncertainty of the ratio is defined in in Section G.1.

Figure 9 shows the ratio of the moles identified by the TCD and MS to the mole injected into the GC/MS as a function of z^* . In most cases, the ratio is close to unity, indicating that the total moles injected into the GC/MS are all accounted for in the calculation. For low values of z^* , the ratio is higher than unity, indicating an error in the calibration which is most

likely due to an over-prediction of the fuel species that are found at high concentrations near the liquid surface.

4.2 Carbon to Hydrogen Ratio

Another way to verify the accuracy of the gas species measurements is to calculate the ratio of carbon to hydrogen atoms contained in all gas species at each vertical measurement location:

$$\frac{C}{H} = \frac{W_C}{W_H} \frac{\sum x_i \bar{X}_i}{\sum y_i \bar{X}_i}$$
(15)

where the summation is over all measured gas species, and x_i and y_i are the numbers of carbon and hydrogen atoms in the molecule, respectively. The carbon to hydrogen ratio of the fuel molecules are reported in Table 1, and the ratio for each fuel is shown in Fig. 10.



Fig. 10. Carbon to hydrogen ratio calculated from all measured gas species compared to the theoretical values. The uncertainty of the ratio is defined in in Section G.2.

A similar verification test is to apply the summations in Eq. (15) to all *product* species; that is, all species except the original fuel molecule and any detected compound that shares the same carbon to hydrogen ratio as the parent fuel. The carbon to hydrogen ratio for the product gases should be the same as that for all species. Figure 11 shows the results. Note that an exact match is not expected because the idealized values do not take into account molecular diffusion, nor the mass of soot, which represents a very small correction.

Another verification test considers the inert species argon and nitrogen whose molar ratio should be the same regardless of fuel type. The measured Ar/N₂ ratio in ambient air is 0.0117 ± 0.0005 , and the measured ratio at different vertical locations within the fires



Fig. 11. Carbon to hydrogen ratio calculated from all measured *product* species compared to the theoretical values. The uncertainty of the ratio is defined in in Section G.2.

is depicted in Fig. 12. All points fall within the uncertainty bounds of the ambient air measurement.



Fig. 12. Ar/N_2 molar ratio at different elevations compared to the uncertainty bounds of the measurement in ambient air. The uncertainty of the ratio is defined in in Section G.3.

4.3 Plotting the Results in Mixture Fraction Space

The mixture fraction, Z, is defined as the mass fraction of the gases containing carbon, in addition to soot, that originate in the fuel stream. It can be expressed³ as follows:

$$Z = \bar{Y}_{\rm F} + \frac{W_{\rm F}}{\rm x} \sum_{i \neq \rm F} \frac{\bar{Y}_i}{W_i} \tag{16}$$

where \bar{Y}_F , W_F , and x are the mass fraction, molecular weight, and number of carbon atoms in the fuel molecule, respectively. Assuming ideal (i.e. no CO or soot), infinitely-fast (fuel and oxygen from the air cannot co-exist) combustion, the mass fractions of all species can be expressed as piece-wise linear "state relations" according to the following reaction:

$$C_{x}H_{y}O_{z} + \eta \left(x + \frac{y}{4} - \frac{z}{2}\right) \left(O_{2} + 3.76 N_{2} + 0.0445 \text{ Ar}\right) \rightarrow \\ \max(0, 1 - \eta) C_{x}H_{y}O_{z} + \max(0, 1 - \eta) \left(x + \frac{y}{4} - \frac{z}{2}\right) O_{2} + \min(1, \eta) x CO_{2} + \\ \min(1, \eta) \frac{y}{2} H_{2}O + \eta \left(x + \frac{y}{4} - \frac{z}{2}\right) \left(3.76 N_{2} + 0.0445 \text{ Ar}\right)$$
(17)

The parameter η is the reciprocal of the local fuel equivalence ratio, ϕ ,

$$\phi = \frac{(F/A)}{(F/A)_{st}} = \frac{1}{\eta}$$
(18)

where F/A is the fuel-air mass ratio and the subscript st denotes the stoichiometric condition.

Figures 13, 14, 15, 16 show the mean mass fraction measurements as a function of the mixture fraction for the methanol, ethanol, acetone, and methane fires, respectively. The dotted lines represent ideal combustion calculated from Eq. (17). The vertical dotted line identifies the stoichiometric value of the mixture fraction. Additional plots that detail the averaged mass fractions of all detected species as a function of mixture fraction, with their expanded uncertainties, are provided in Appendix F.

Where the mixture fraction is much less than stoichiometric, all major gas species are in close agreement with the ideal state relations; the measured mass fractions of unburned fuel and CO are nearly zero, and the O_2 is close to its respective theoretical value. The measured mass fraction of CO_2 and H_2O are found to peak close to the stoichiometric mixture fraction. In the fuel-rich region, the measured mass fraction of CO_2 differs considerably from the ideal state relation due to the substantial portion of CO and soot.

³The uncertainty of the mixture fraction is determined from propagating the error in the mass fraction measurements. A detailed description of the mixture fraction uncertainty is provided in Appendix G.4.



Fig. 13. Mean mass fractions as a function of mixture fraction, methanol. The uncertainty of the mean mass fractions shown here is defined in Section B.2.



Fig. 14. Mean mass fractions as a function of mixture fraction, ethanol. The uncertainty of the mean mass fractions shown here is defined in Section B.2.



Fig. 15. Mean mass fractions as a function of mixture fraction, acetone. The uncertainty of the mean mass fractions shown here is defined in Section B.2.



Fig. 16. Mean mass fractions as a function of mixture fraction, methane. The uncertainty of the mean mass fractions shown here is defined in Section B.2.

5. Conclusion

In summary, time-averaged local measurements of temperature and gas species concentrations are made to characterize the structure of methanol, ethanol, and acetone 30 cm diameter pool fires and a methane 38 cm pool fire steadily burning in a quiescent environment. A verification scheme is developed to verify the gas species measurements that considered the accuracy of the calibration and the overall stoichiometry of combustion for each fuel. The gas species measurements are favorably compared to the idealized SCR values, which lends confidence to the veracity of the measurements. These local measurements complement previous measurements and provide insight into the complex chemical structure of medium-scale pool fires.

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A. Uncertainty Analysis of Pool Fire Parameters

A.1 Mass Burning Rate

The mean mass burning rate, \dot{m} , is determining by weighing the fuel reservoir at the start and end of the steady burning stage using a Precisa XB-6200C Precision balance, calibrated using a collection of standard weights. The Type A evaluation of uncertainty is the standard deviation in the measurements made during replicate experiments, $s_{\dot{m}}$. The Type B evaluation of uncertainty is the reported accuracy of the balance, 0.01 g, divided by the time interval. The Type A uncertainty dominates; thus, the standard uncertainty is approximately the standard deviation of the multiple measurements:

$$u_{in} \approx s_{in}$$
 (A.1)

A.2 Heat Release Rate and Q^*

The heat release rate, \dot{Q} , is the product of the fuel mass loss rate, \dot{m} , and the heat of combustion, ΔH . The relative uncertainty of the former is far greater than the latter, thus:

$$u_{\dot{Q}} \approx \Delta H \, u_{\dot{m}} \tag{A.2}$$

The uncertainty of Q^* , is dominated by the relative uncertainty of \dot{Q} , thus:

$$u_{Q^*} \approx \frac{u_{\dot{Q}}}{c_p \rho_\infty T_\infty \sqrt{g} D^{5/2}} \tag{A.3}$$

A.3 Mean Flame Height

The mean flame height, $L_{\rm f}$, is determined using photographic analysis as described in Section 2. The Type A evaluation of uncertainty is the standard deviation of the height measurements, $s_{L_{\rm f}}$, made for each frame. The uncertainty of the distance measurement for each frame is relatively small, 0.1 %, and thus the standard uncertainty is approximately the standard deviation of the heights over multiple frames:

$$u_{L_{\rm f}} \approx s_{L_{\rm f}}$$
 (A.4)

A.4 Uncertainty Analysis of Temperature Measurements

The uncertainty of the temperature measurements is estimated from Type A and Type B uncertainty. The Type A evaluation of uncertainty is the standard deviation of the corrected temperature measurements made at each flame location. The bias error sources, Type B uncertainty, in the S-type thermocouple, u_{inst} , used to measure temperature (1.5 °C) of the gas sample injected is relatively small compared to the Type A uncertainty.

B. Uncertainty Analysis of Gas Species Concentrations

B.1 Uncertainty of Volume Fractions

As shown in Eq. 7, volume fraction, \bar{X}_i , is calculated from the ratio between the number of moles of a given species, n_i , and the total number of moles identified, n_{tot} . The uncertainty of the measured volume fraction is estimated using the law of propagation of uncertainty after determining the volume fraction of each species:

$$u_{\bar{X}_{i}} = \sqrt{\left(\frac{\partial \bar{X}_{i}}{\partial n_{i}} u_{n_{i}}\right)^{2} + \left(\frac{\partial \bar{X}_{i}}{\partial n_{\text{tot}}} u_{n_{\text{tot}}}\right)^{2}}$$
(B.1)

A coverage factor of 2 is applied to the combined uncertainty to produce a 95 % confidence interval.

B.1.1 Number of Moles of a Given Species

The number of moles of a given species is determined from a calibration function of the integrated peak area of the respective species obtained from the TCD's and the MS's Total Ion Current (TIC) chromatograms. The Type A evaluation of standard uncertainty of the number of moles of a given species is taken as the standard deviation of the measurements obtained from the repeated tests. The Type B evaluation of standard uncertainty is determined from the error in the calibration functions for each species measured by the TCD and MS is further detailed in Appendix C. The combined uncertainty is found via quadrature:

$$u_{n_i} = \sqrt{u_{n_{i,\text{cal}}}^2 + s_{n_i}^2} \tag{B.2}$$

B.1.2 Total Number of Moles Identified

The total number of moles detected is determined from the summation of the number of moles for each species identified by the TCD and TIC chromatograms. Therefore, the uncertainty in the total number of moles identified is the combined uncertainty of all the identified species via quadrature:

$$u_{n_{\text{tot}}} = \sqrt{\sum_{n=1}^{N} s_{n_i}^2} \tag{B.3}$$

where N is the number of a species identified species in the TCD and TIC chromatogram.

B.2 Uncertainty of Mass Fractions

Mass fraction of a given species is calculated using its measured volume fraction, \bar{X}_i , molecular weight, W_i , and the average molecular weight of all detected gas species, W_{tot} , as

shown in Eq. 8. The uncertainty of the mass fraction of a given species is estimated from the law of propagation of uncertainty using Eq. 8.

$$u_{\bar{Y}_{i}} = \sqrt{\left(\frac{\partial \bar{Y}_{i}}{\partial \bar{X}_{i}} u_{\bar{X}_{i}}\right)^{2} + \left(\frac{\partial \bar{Y}_{i}}{\partial W_{\text{tot}}} u_{W_{tot}}\right)^{2}}$$
(B.4)

B.2.1 Uncertainty of the average molecular weight

The uncertainty of the average molecular weight is determined using the law of propagation of uncertainty, which accounted for each detected species measured from the injected sample.

$$u_{W_{tot}} = \sqrt{\left(\sum u_{\bar{X}_i} W_i\right)^2} \tag{B.5}$$

Uncertainty Analysis of Gas Species Calibrations

С.

$$n_i = a \,(\text{Area}) + \mathbf{b} \tag{C.1}$$

The coefficients in the calibration curves are weighted to account for the error of each gas standard used in the calibration procedure. The uncertainty of a number of moles determined from calibration function is estimated using the law of propagation of uncertainty:

$$u_{n_{i}} = \sqrt{\left(\frac{\partial n_{i}}{\partial a}u_{a}\right)^{2} + \left(\frac{\partial n_{i}}{\partial b}u_{b}\right)^{2}}$$
(C.2)

The uncertainties of the slope and intercept in a weighting linear regression are as follows:

$$u_{\rm a} = \sqrt{\frac{\sum \frac{1}{u_{n_{i,\rm cal}}}}{(\sum \frac{\mathrm{Area}_{\rm i}^2}{u_{n_{i,\rm cal}}^2})(\sum \frac{1}{u_{n_{i,\rm cal}}^2}) - (\sum \frac{\mathrm{Area}_{\rm i}}{u_{n_{i,\rm cal}}^2})^2}}$$
(C.3)

$$u_{\rm b} = \sqrt{\frac{\sum \frac{\operatorname{Area}_{\rm i}^2}{u_{n_{i,\rm cal}}}}{(\sum \frac{\operatorname{Area}_{\rm i}^2}{u_{n_{i,\rm cal}}^2})(\sum \frac{1}{u_{n_{i,\rm cal}}^2}) - (\sum \frac{\operatorname{Area}_{\rm i}}{u_{n_{i,\rm cal}}^2})^2}$$
(C.4)

where $u_{n_{i,cal}}$ is the uncertainty of a known number of moles injected into the GC/MS for calibration.

During calibration, the number of moles of a given species $n_{i,cal}$ are calculated from the product of the total moles injected into the GC/MS, n_{inj} , and the known concentration of the particular species in the calibration standard, C_i .

$$n_{i,\text{cal}} = C_i n_{\text{inj}} \tag{C.5}$$

A collection of gas calibration standards for a variety of species are pre-selected to provide a broad range of concentrations. All calibration standards are mixtures of the target gas species with a nitrogen balance, with the exception of one standard balanced in Air. A list of gas standards used in this work, with their respective concentrations and Type B evaluation of standard uncertainty, is provided in Appendix C.2.

The uncertainty of the number of moles of a given species injected into the GC/MS for calibration is estimated using the law of propagation of uncertainty:

$$u_{n_{i,cal}} = \sqrt{\left(\frac{\partial n_{i,cal}}{\partial C_i} u_{C_i}\right)^2 + \left(\frac{\partial n_{i,cal}}{\partial n_{inj}} u_{n_{inj}}\right)^2}$$
(C.6)

C.1 Total Moles Injected into the GC/MS for Calibation

The total moles injected into the GC/MS, n_{inj} , for calibration is determined from Eq. (14) using the pressure, P, temperature, T, and volume, V_s , of the gas sample injected into the GC/MS. Pressure and temperature measurements are made using a digital pressure gauge (OMEGA P10R-001) and K-type thermocouple located at the GC/MS sample loop injection valve, respectively, sampling at 2 Hz for 50 s. The volume of the GC/MS sample loop is 2 mL. The Type A evaluation of uncertainty of the total moles injected into the GC/MS for calibration is determined from the standard error of the pressure, s_P , and temperature, s_T readings from the sampling period. The Type B evaluation of uncertainty for the total moles injected into the GC/MS for calibration, u_{inst} , used to measure pressure (0.008% accuracy of the reading) and temperature (1.5 °C) of gas sample injected. The combined uncertainty of the temperature measurements is found via quadrature:

$$u_T = \sqrt{u_{\text{inst}}^2 + s_T^2} \tag{C.7}$$

The Type A uncertainty is the dominant contribution for the uncertainty of the pressure measurements, therefore its uncertainty of pressure is approximately the standard deviation.

$$u_P \approx s_P^2$$
 (C.8)

The standard uncertainty of the total moles injected into the GC/MS for calibration is estimated using the law of propagation of uncertainty:

$$u_{n_{\rm inj}} = \sqrt{\left(\frac{\partial n_{\rm inj}}{\partial P} u_P\right)^2 + \left(\frac{\partial n_{\rm inj}}{\partial T} u_T\right)^2} \tag{C.9}$$

C.2 Table of Gas Standards with Error

A table of the gas standards with their respective concentrations and Type B evaluation of standard uncertainty, used for calibrating the GC/MS is provided below. Lot numbers for all standards are provided for traceability.

Components	Uncertainty(%)	Distributor	Lot No.
200 ppm Acetone	2.00	Gasco Affiliates, LLC.	DNJ-ACE-200N-1
0.26% Acetylene	2.00	Gasco Affiliates, LLC.	FBJ-M24-0.25%-1
1.04% Acetylene	2.00	Gasco Affiliates, LLC.	FBJ-M24-1
1.02% Argon	2.00	Gasco Affiliates, LLC.	DBJ-2-1N-1
88.5% Argon	2.00	Gasco Affiliates, LLC.	DBJ-2-90N-1
100 ppm Benzene	2.00	Gasco Affiliates, LLC.	FBJ-21-100-3
15.6% Carbon Dioxide	0.04	NIST Gas Sensing Metrology Group	9-C-44
24.5% Carbon Dioxide	2.00	Gasco Affiliates, LLC.	KBI-35-25-1
1.00% Carbon Dioxide	2.00	Matheson Tri-Gas	9306620888
2.51% Carbon Dioxide	2.00	Roberts Oxygen	1002080917
7.00% Carbon Dioxide	2.00	Roberts Oxygen	1009010318
9.00% Carbon Dioxide	2.00	Praxair Doistribution Inc.	304113044702
0.30% Carbon Monoxide	2.00	Roberts Oxygen	1009010318
0.02% Carbon Monoxide	2.00	Matheson Tri-Gas	9306620888
0.11% Carbon Monoxide	2.00	Roberts Oxygen	1002080917
4.00% Carbon Monoxide	2.00	Praxair Doistribution Inc.	304113044702
7.81% Carbon Monoxide	0.02	NIST Gas Sensing Metrology Group	51-28-C
0.51% Ethane	2.00	Gasco Affiliates, LLC.	FBJ-62N-0.5-1
1.00% Ethane	2.00	Gasco Affiliates, LLC.	FBJ-152N-1-1%-1
2.55% Ethane	2.00	Gasco Affiliates, LLC.	FBJ-152N-2.5-1
0.51% Ethylene	2.00	Gasco Affiliates, LLC.	FBJ-62N-0.5%-1
1.02% Ethylene	2.00	Gasco Affiliates, LLC.	FBJ-62N-1%-1
2.55% Ethylene	2.00	Gasco Affiliates, LLC.	FBJ-62N-2.5%-1
0.26% Hydrogen	2.00	Gasco Affiliates, LLC.	FBJ-84-0.25-1
0.50% Hydrogen	2.00	Gasco Affiliates, LLC.	KBI-84-0.5-1
1.00% Hydrogen	2.00	Gasco Affiliates, LLC.	KBI-84-1-3
2.00% Hydrogen	2.00	Gasco Affiliates, LLC.	FBJ-84-2-5
4.03% Hydrogen	2.00	Gasco Affiliates, LLC.	FBJ-84-4-2
0.40% Methane	2.00	Gasco Affiliates, LLC.	DBJ-135N-0.4-1
3.95% Methane	2.00	Gasco Affiliates, LLC.	DBJ-135N-4-2
40.8% Methane	2.00	Gasco Affiliates, LLC.	DBJ-135N-40-1
0.50% Oxygen	2.00	Gasco Affiliates, LLC.	DBJ-2-90N-1
1.97% Oxygen	0.01	NIST Gas Sensing Metrology Group	73-D-03
5.02% Oxygen	2.00	Gasco Affiliates, LLC.	DBJ-161-5-5
9.92% Oxygen	0.02	NIST Gas Sensing Metrology Group	72-D-60
10.2% Oxygen	2.00	Gasco Affiliates, LLC.	KBI-161-10-6
20.7% Oxygen	0.04	NIST Gas Sensing Metrology Group	71-D-51
0.42% Propane	2.00	Gasco Affiliates, LLC.	DBJ-176N-0.4-1
39.6% Propane	2.00	Gasco Affiliates, LLC.	DBJ-176N-40-1

Table 2. Gas standards used to calibrate the GC/MS

C.3 Concentration of Vapors from Bubblers

The volume fraction of vapor from liquid materials are calibrated from the ratio of the liquid-vapor pressure in the heated flask to the total pressure in the flask.

$$C_{\rm vap} = \frac{P_{\rm vap}}{P} \tag{C.10}$$

The bubbler setup shown in Fig. 17:



Fig. 17. Flow diagram for bubble calibration system used for liquid materials (acetone, ethanol, methanol, and water)

The material of interest is placed in a 500 ml Pyrex flask sitting on a heating plate. Nitrogen, acting as a carrier gas, is bubbled through the liquid bath and then transported through a heated gas line and into the GC/MS sample loop. Vapor from the liquid is transported along with the carrier gas in an amount liquid depending on the temperature. The concentration of the vapor injected into the GC/MS is calculated from a liquid-vapor pressure correlation provided by DIPPR [9].

$$P_{\rm vap} = e^{A + \frac{B}{T_{\rm B}} + C \ln(T_{\rm B}) + D(T_{\rm B})^E}$$
(C.11)

In this correlation, P_{vap} is the vapor pressure calculated from the temperature of the liquid bath, T_{B} , with the coefficients (A, B, C, D, E) specific to the liquid material. Table 3 lists all coefficients for each calibrated liquid, including the uncertainty of their respective correlations, u_{corr} .

Liquid Material	Α	В	С	D	Е	Uncertainty(%)
Acetone	69.006	-5599.6	-7.0985	6.2237E-6	2.00	3.00
Ethanol	73.304	-7122.3	-7.1424	2.8853E-6	2.00	1.00
Methanol	82.718	-6904.5	-8.8622	7.4664E-6	2.00	3.00
Water	73.649	-7258.2	-7.3037	4.1653E-6	2.00	0.20

Table 3. Liquid vapor pressure correlation coefficients for various calibrated liquids

The concentration range of each calibrated liquid is approximately 2 % to 50 %. Liquid bath temperatures are controlled using a heating plate positioned underneath the insulated bubbler. The temperature of the bath is measured using a K-type thermocouple placed at the liquid surface. The bath temperature measurements are sampled at 2 Hz for 50 s simultaneously with pressure and temperature measurements of the GC/MS sample loop. Liquid-vapor calibrations are conducted once the bath reaches a steady-state temperature (approximately 1 h) and the nitrogen/vapor gas mixture has swept through the sample loop. Upon injection into the GC/MS, pressure and temperature measurements of the sample loop. Ioop are made as described in Appendix C.1.

The uncertainty of the concentration determined using Eq. (C.10) is estimated using the law of propagation of uncertainty:

$$u_{C_{\text{vap}}} = \sqrt{\left(\frac{\partial C_{\text{vap}}}{\partial P} u_P\right)^2 + \left(\frac{\partial C_{\text{vap}}}{\partial P_{\text{vap}}} u_{P_{\text{vap}}}\right)^2}$$
(C.12)

The uncertainty of the pressure measured upon injected is calculated from Eq. (C.8). The uncertainty of the vapor pressure is found by combining the propagated error of liquid bath temperature and the uncertainty in the correlation via quadrature:

$$u_{P_{\text{vap}}} = \sqrt{\left(\frac{\partial P_{\text{vap}}}{\partial T_{\text{B}}} u_{T_{\text{B}}}\right)^2 + u_{\text{corr}}^2}$$
(C.13)

The Type A evaluation of uncertainty of the liquid bath temperature readings is determined from the standard error of the temperature, s_{T_B} readings from the sampling period. The Type B evaluation of uncertainty for the liquid bath temperature is defined as the bias error source (1.5 °C) in the thermocouple, u_{inst} . The combined uncertainty liquid bath temperature is determined via quadrature:

$$u_{T_{\rm B}} = \sqrt{u_{\rm inst}^2 + s_{T_{\rm B}}^2}$$
 (C.14)

D. Uncertainty Analysis of the Soot Mass Fraction

The local soot mass fraction measurements, Y_s , made at various heights above the fuel surface are calculated through a combination of Eqs. (9), (10), and (11):

$$Y_{\rm s} = \frac{m_{\rm s} V_s}{\dot{V} \Delta t \, m_{\rm tot}} \frac{T_{\infty}}{T_{\rm g}} \tag{D.1}$$

where m_s is the mass of soot collected on the PTFE filter and gun cleaning patches, \dot{V} is the volumetric flow rate, V_s is the volume of the sample loop, m_{tot} is the total mass of the gas sample detected in the TCD and TIC chromatograms, T_{∞}/T_g is the ratio of the internal gas flow temperature readings of the mass flow controller to the temperature of the probe, and Δt is the total sampling time. The uncertainty of the measured soot mass fraction is estimated using the law of propagation of uncertainty after determining the soot mass fraction:

$$u_{Y_{\rm s}} = \sqrt{\left(\frac{\partial Y_{\rm s}}{\partial m_{\rm s}} \, u_{m_{\rm s}}\right)^2 + \left(\frac{\partial Y_{\rm s}}{\partial \dot{V}} \, u_{\dot{V}}\right)^2 + \left(\frac{\partial Y_{\rm s}}{\partial m_{\rm tot}} \, u_{m_{\rm tot}}\right)^2} \tag{D.2}$$

The uncertainty of the temperature measurements is not accounted for since the uncertainty of the soot mass fraction is dominated by the remaining parameters uncertainties.

D.1 Mass of Soot

The mass of soot is determined from the difference in mass of the dried PTFE filter and gun cleaning patches before and 48 h after each experiment. The Type A evaluation of standard uncertainty of the mass of soot, m_s , is taken as the standard deviation, s_{m_s} , of the measurements sampled three times before and after each test. The Type B evaluation of uncertainty, u_{inst} , is determined from the instrumentation error sources of the scale and is found to be 1 % of the reading. The Type A evaluation of uncertainty dominates; thus, the standard uncertainty is approximately the standard deviation of the multiple measurements:

$$u_{m_{\rm S}} \approx s_{m_{\rm S}}$$
 (D.3)

D.2 Mass Flow Controller Volumetric Flow Rate

A mass flow controller is used to measure the volumetric flow rate, \dot{V} , within the gas sampling line. The Type A uncertainty is taken as the standard deviation of the flow measurements sampled at 2 Hz during the sampling period which varies from 12 min to 25 min depending on the sampling location within the fire. The Type B sources of uncertainty consist of the calibration error, u_{cal} , and the precision error sources at calibration conditions, u_{prec} . The calibration error is given as 2 ml/min. The precision error is 0.8 % of the reading plus 0.2 % of the full scale (2 L/min). The combined uncertainty is calculated via quadrature:

$$u_{\dot{v}} = \sqrt{u_{\text{prec}}^2 + u_{\text{cal}}^2 + s_{\dot{v}}^2}$$
 (D.4)

D.3 Total Mass Identified

The total mass detected in the TCD and TIC chromatograms, m_{tot} , is calculated:

$$m_{\text{tot}} = \sum_{n=1}^{N} n_i \mathbf{W}_i \tag{D.5}$$

where n_i is the moles of species *i* and W_i is the molar mass. The uncertainty in the total mass is calculated from the uncertainties of all identified species, defined in Section B.1.2, multiplied by their corresponding molar mass via quadrature:

$$u_{m_{\text{tot}}} = \sqrt{\sum_{n=1}^{N} (u_{n_i} \mathbf{W}_i)^2}$$
(D.6)

where *N* is the number of a species identified species in the chromatograms.

E. Figures of Averaged Volume Fractions

E.1 Methanol



Fig. 18. Plot of volume fractions of all species identified in the methanol pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.

E.2 Ethanol



Fig. 19. Plot of volume fractions of major species identified in the ethanol pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.



Fig. 20. Plot of volume fractions of minor and trace species identified in the ethanol pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.

E.3 Acetone



Fig. 21. Plot of volume fractions of major species identified in the acetone pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.



Fig. 22. Plot of volume fractions of minor and trace species identified in the acetone pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.

E.4 Methane



Fig. 23. Plot of volume fractions of all species identified in the methane pool fire as a function of z^* along the pool centerline. The error is a combined uncertainty, further described in Section B.



Fig. 24. Plot of volume fractions of minor and trace species identified in the methane pool fire as a function of z^* along the pool centerline. The error presented here is a combined uncertainty, further described in Section B.

F. Figures of Averaged Mass Fractions

F.1 Methanol



Fig. 25. Plot of mass fractions, with uncertainty, of major species identified in the methanol pool fire centerline as function of mixture fraction. The uncertainty is a combined uncertainty, discussed in further detail in Sections G and B for the mass fraction and mixture fractions, respectively.

F.2 Ethanol



Fig. 26. Plot of mass fractions, with uncertainty, of major species identified in the ethanol pool fire centerline as function of mixture fraction. The uncertainty is a combined uncertainty, discussed in further detail in Sections G and B for the mass fraction and mixture fractions, respectively.

F.3 Acetone



Fig. 27. Plot of mass fractions, with uncertainty, of major species identified in the acetone pool fire centerline as function of mixture fraction. The uncertainty is a combined uncertainty, discussed in further detail in Sections G and B for the mass fraction and mixture fractions, respectively.

F.4 Methane



Fig. 28. Plot of mass fractions, with uncertainty, of major species identified in the methane pool fire centerline as function of mixture fraction. The uncertainty is a combined uncertainty, discussed in further detail in Sections G and B for the mass fraction and mixture fractions, respectively.

G. Uncertainty Analysis of the Verification Scheme

G.1 Ratio of moles identified to moles injected

The uncertainty of the ratio of the total number of moles identified by the TCD and MS, n_{tot} , to the number of moles injected into the GC/MS, n_{inj} , is calculated using the law of propagation of uncertainty and shown in the equation below:

$$u_{n_{\text{tot}}/n_{\text{inj}}} = \sqrt{\left(\frac{u_{n_{\text{tot}}}}{n_{\text{inj}}}\right)^2 + \left(\frac{-n_{\text{tot}} u_{n_{\text{inj}}}}{n_{\text{inj}}^2}\right)^2} \tag{G.1}$$

The uncertainty of the total number of moles identified, n_{tot} , is described in Section B.1.2. The uncertainty of the number of moles injected into the GC/MS, n_{inj} , is described in Section C.1.

G.2 Carbon to Hydrogen Ratio

The carbon to hydrogen ratio is calculated using Eq. 15 using the mole fraction of any quantified gas species that contained either carbon or hydrogen.

$$u_{\rm C/H} = \frac{W_{\rm C}}{W_{\rm H}} \sqrt{\sum \left(\frac{\partial ({\rm C/H})}{\partial \bar{X}_i} \, u_{\bar{X}_i}\right)^2 + \sum \left(\frac{\partial ({\rm C/H})}{\partial X_i} \, u_{\bar{X}_i}\right)^2} \tag{G.2}$$

The uncertainty of the mass fractions of carbon and hydrogen containing species is described in Appendix B.2.

G.3 Inert Ratio

The inert ratio is calculated from the ratio the Nitrogen and Argon volume fractions, \bar{X}_{N_2} and \bar{X}_{Ar} , identified by the TCD and TIC chromatograms. The uncertainty of inert ratio is determined from the law of propagation of uncertainty:

$$u_{\bar{X}_{Ar}/\bar{X}_{N_{2}}} = \sqrt{\left(\frac{u_{\bar{X}_{Ar}}}{\bar{X}_{N_{2}}}\right)^{2} + \left(\frac{-\bar{X}_{Ar}u_{\bar{X}_{N_{2}}}}{\bar{X}_{N_{2}}}\right)^{2}}$$
(G.3)

The uncertainty of the volume fractions is described in Appendix B.1.

G.4 Uncertainty Analysis of the Mixture Fraction

The mixture fraction, Z, is determined from Eq. (16). The uncertainty of the mixture fraction is a function of the uncertainties in the carbon carrying species:

$$u_{Z} = \sqrt{u_{\bar{Y}_{F}}^{2} + \left(\frac{W_{F}}{x}\right)^{2} \sum_{i \neq F} \left(\frac{u_{\bar{Y}_{i}}}{W_{i}}\right)^{2}} \tag{G.4}$$

The uncertainty of the mass fractions of carbon containing species is discussed in Appendix B.2.

H. Previous Measurements in 30 cm Pool Fires

Many experiments have been conducted on pool fires such as those considered in this report [2–4, 7, 8, 10, 16, 22–26]. Table 4 lists previously reported local and global measurements characterizing the structure of 30 cm diameter methanol, ethanol and acetone pool fires steadily burning under well-ventilated and quiescent conditions. These measurements complement each other and help build a more complete picture of the energetics, structure and dynamics of medium-scale pool fires burning oxygenated fuels. The accumulated information provides a basis for understanding details about these fires that makes them particularly suitable candidates for fire model evaluation. The three fires considered here are particularly useful as a testbed for radiation sub-models since blackbody radiation from soot is dependent on the fuel type. The time-averaged mean and RMS data from several of the studies is available through the Measurement and Computational Fire Phenomena (MaCFP) GitHub website [27].

Table 4. List of previously measured parameters obtained from a well-ventilated, round, steady,30 cm diameter pool fires burning in a quiescent environment.

	Parameter	References			
		Methanol	Ethanol	Acetone	
	Mass Loss Rate	[4]		[24]	
al	Heat Release Rate	[8]		[25]	
lob	Mean Flame Height	[4, 7, 22, 23]	[4, 7, 23]	[7, 23, 24]	
9	Pulsation Frequency	[2-4, 7, 8, 22, 26]	[7, 23, 26]	[7, 23, 25, 26]	
	Radiative Fraction	[7]	[7]	[7]	
	Radiative flux distribution on fuel surface	[3]			
	Total Heat Flux Distributuion on Fuel Surface	[7]	[7]	[7]	
al	Radiative flux to surroundings	[4, 7, 22]			
Loc	Transient fuel temperature	[3]			
	Gas species volume fraction	[2]			
	Local Gas Phase Temperature	[4, 8]			
	Local Gas Phase Velocity	[8]			

There are subtle differences in the experimental conditions as well as reporting assumptions and approximations that may account for some of the differences among the reported results. Some of the relevant issues are highlighted in Table 5 below for the methanol, ethanol and acetone pool fires, respectively. The tables include information on the steady-state mass burning flux, m'', lip height, mean flame height.

The radiative fraction to the surroundings including the fuel pool, χ_{rad} , and excluding the fuel pool, χ_r , during steady burning is reported for each study. χ_{rad} is defined as the total radiative heat transfer to the surroundings and onto the fuel surface such that:

$$\chi_{\rm rad} = \chi_{\rm r} + \chi_{\rm sr} \tag{H.1}$$

where χ_r is the integrated radiative flux emitted by the fire in all directions except to the fuel surface, normalized by the fire heat release rate. The term χ_{sr} represents the integrated

radiative flux emitted by the fire towards the fuel surface, normalized by the fire heat release rate. The values of the fractional enthalpy terms (χ_r , χ_{sr}) vary from fire to fire and are dependent on fuel type, burner diameter, and fire size. Ref. [7] describes the development of these terms in more detail.

There are several general considerations regarding the previous work:

- Ref. [8] reports using a round, stainless-steel, 30.5 cm diameter burner, which differs from the 30.1 cm (inner) diameter burner used in Refs. [2–4, 7, 10, 22, 23]. Due to this disparity (3 % in area), Table 4 refers to the mass burning flux (rather than the mass burning rate). The same circular, water-cooled, stainless-steel pan was used for experiments [2–4, 7, 10, 22, 23]. The burner had an inner diameter of 30.1 cm, a depth of 15 cm, and a wall thickness of 1.6 mm. The burner was fitted with legs such that the burner rim was positioned 30 c m above the ground. The bottom of the burner was maintained at a constant temperature by flowing tap water (nominally 20 °C) through a 3 cm section on the bottom of the fuel pan. The exhaust flow in [7] was 0.50 kg/s; this value can be assumed to be approximately the same for Refs. [2, 3, 10, 22, 23].
- The lip height reported in Ref. [8] and several of the other studies was 10 mm [2, 10], except Refs. [2–4, 22, 23] where it was reported as 4 mm or 5 mm (see Table 4).
- Weckman and Strong's results [8] and Ref. [2] are presented as a function of distance from the fuel surface, whereas all the other studies use the top of the burner rim as the point of reference.
- A water-cooling section was not reported on the bottom of the burner used in Ref. [8]; the burner in Refs. [2–4, 7, 10, 22, 23] had a water-cooling section on the bottom of the burner.
- Refs. [7, 23] reported radiative fraction emitted to the surroundings and did not include flux incident on the fuel surface; Ref. [7] reported radiative fraction to the surroundings plus flux incident on the fuel surface; Ref. [22] assumed flux incident on the pool was uniform across the pool and equal to its value just outside the pool burner.
- The amount of CO in the methanol, ethanol and acetone exhaust stream was below detection limits [7], so the combustion efficiency can be assumed to be about 1.
- In Ref. [22], the average heat flux to the pool was assumed to be equal to the radiative flux measured just outside the burner, which is smaller than that expected [7]. Here, χ_{rad} is recalculated using $\chi_r = 0.065$ from [7].
- In Ref. [22], χ_{rad} was based on a heat of combustion of 22.317 kJ/g. Here, χ_{rad} is recalculated using the "net" heat of combustion (19.94 kJ/g) in which gaseous water is assumed to be the product of combustion [9].

• The heat of combustion values provided in Table 1 were used to determine the heat release rates reported in Table 5.

	$\dot{m''}$ (g/m ² s)	\dot{Q} (kW)	χr	$\chi_{\rm rad}$	Lip	Ref.	Comments
					Hei	ght	
	15.1	21.4			10	[8]	Volumetric Burning
							Rate of 1.35 cm^3/s
	13.5 ± 1.2	19.2	0.19	0.24	10	[7]	
_	12.9	18.3	0.22		5	[3]	Fuel Mass Burning
ou							Rate of 0.92 g/s
tha	12.6	18.0		0.20	5	[22]	Fuel Mass Burning
Me							Rate of 0.9 g/s
		20	0.18	0.22	4	[4]	
					5	[2]	
					10	[10]	
	12.8 ± 1.3	18.2	0.22		5	[23]	
	12.4 ± 0.5	17.6			10	This study	
	12.4 ± 0.5	17.6			10	This study	
	13.2 ± 0.4	18.7 ± 0.6	0.20 ± 0.02	0.22 ± 0.2		Average	
Ы	16.2 ± 1.8	31.0	0.21		10	[7]	
anc	14.4 ± 1.4	27.3	0.16		5	[23]	
∃th	13.9 ± 2.1	26.3		0.26	10	This study	
	14.8 ± 1.2	28.2 ± 2.5	0.19 ± 0.3	0.26		Average	
	25.9	54.1	0.26		10	[8]	$\dot{Q}^{\prime\prime}$ of 740 kW/m ²
one	18.7 ± 2.1	38.1	0.27	0.31	10	[7]	
cetc	18.7 ± 1.9	37.7	0.28		5	[23]	
Αc	17.6 ± 2.6	35.5			10	This Study	
	$18.3^{**} \pm 0.6$	37.1 ± 1.4	0.27 ± 0.1	0.31		Average	

Table 5. Reported mass burning flux, m'', heat release rate, \dot{Q} , radiative fractions, excluding, χ_r , and including, χ_{rad} the fuel pool, and the lip height of steady burning pool fires

** The mass burning flux of acetone reported in [8] is an outlier and not included in the average. Uncertainty of averages is defined as the standard deviation with a coverage factor of 2.