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Technical Note

No. 333

**THE VISCOSITY AND THERMAL CONDUCTIVITY
COEFFICIENTS OF DILUTE ARGON BETWEEN
100 AND 2000°K**

H. J. M. HANLEY



U. S. DEPARTMENT OF COMMERCE
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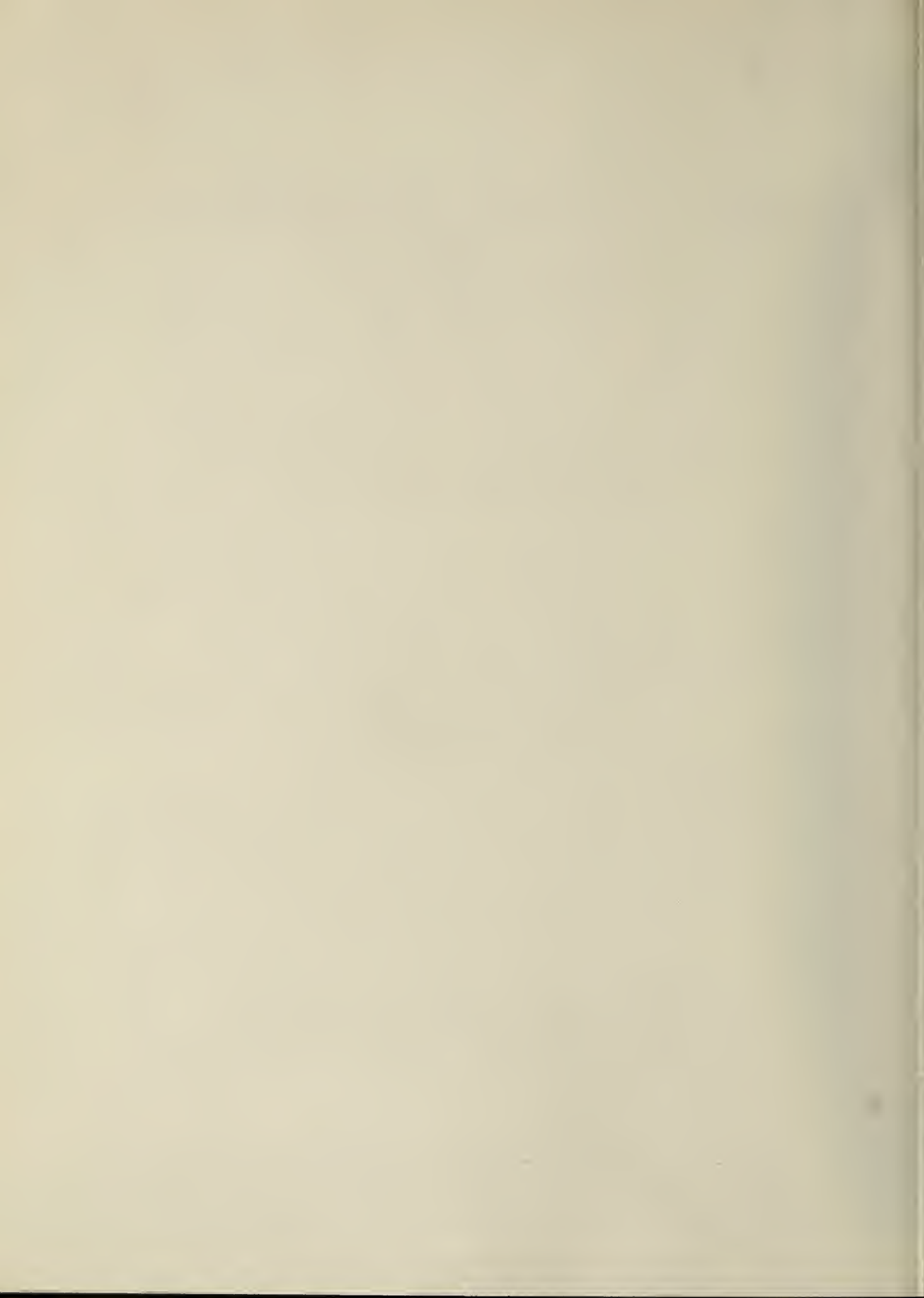
Technical Note 333

ISSUED March 11, 1966

THE VISCOSITY AND THERMAL CONDUCTIVITY COEFFICIENTS OF DILUTE ARGON BETWEEN 100 AND 2000 °K

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THE VISCOSITY AND THERMAL CONDUCTIVITY COEFFICIENTS OF DILUTE ARGON
BETWEEN 100 AND 2000°K FROM THE KIHARA POTENTIAL FUNCTION

H. J. M. Hanley

The variation of the force constants of the Lennard-Jones, Exp: 6, and Kihara potential functions was investigated by comparing the Chapman-Enskog kinetic theory expression for the viscosity coefficient with the experimental viscosity of dilute argon. It was found that this variation was more pronounced than expected. It was necessary to rationalize the choice of the force constants of each function before using the function to compare theory with experimental data. Of the three, the Kihara was found to give the best correlation, and tables of the viscosity and thermal conductivity coefficients of dilute argon between 100 and 2000°K were computed from this potential and the Chapman-Enskog equations.

Key Words: Force-constants, Potential functions, Lennard-Jones, Kihara, Exp: 6, Viscosity, Argon, Correlation, Thermal conductivity, Self-diffusion, Kinetic theory

1. INTRODUCTION

In this note the Chapman-Enskog kinetic theory expressions for the viscosity, thermal conductivity, and self-diffusion coefficients were correlated with experimental data and tables of the two former properties are presented.

Argon was chosen because it is the simplest molecule to treat theoretically and because of the relatively large amount of experimental data available. Although, of course, there are numerous discussions on this topic, and there are several tables of argon transport properties in the literature [1-5]*, a really satisfactory treatment of even dilute argon has not been achieved. This is not unexpected in view of the difficulties of examining any realistic model in a non-equilibrium system. For instance, not only are there many claims that a particular potential function is to be recommended, but there is also wide disagreement over the recommended values for the force-constants or parameters for a given function, such as the Lennard-Jones, (Table I).

TABLE I

$\epsilon/k, ^\circ K$	$\sigma, \text{ \AA}$	Ref.
93.3	3.542	2
119.1	3.491	3
119.5	3.421	4
128.0	3.398	6
152	3.305	7

Table I. Parameters for the Lennard-Jones function, (Eq. 1) selected by various workers.

In view of this rather unsatisfactory situation it was felt that still another examination of the data was justified. This has turned out to be fruitful in that the potential function was found to vary in a manner not fully discussed previously [6, 8]. It will also be shown that the correlation between theory and experimental data is much more affected by the choice of the force-constants than previously reported. It is hoped that the method given will be suitable for other gases of relatively simple structure.

* Numbers in brackets refer to references.

The selection of a function and a choice of "best values" for the corresponding force-constants was based on an examination of the viscosity coefficient alone. Experimental thermal conductivity and self-diffusion coefficients are not as extensive nor as reliable as desired, and it is well known that potential parameters obtained from sources such as the equation of state, often lead to discrepancies when applied to transport phenomena (this point will be briefly commented on in Appendix III).

2. THE POTENTIAL FUNCTIONS

The discussion was restricted to the three most commonly used functions, the Lennard-Jones, the Exp: 6, and the Kihara. The Kihara, in particular, has received much attention in the literature recently [3, 9, 10]. As the functions are well known and have been fully discussed [1, 11, 12], it is only necessary to outline them here. If $U(r)$ is the interaction potential of two molecules separated by distance r , and ϵ is the maximum energy of attraction, or energy minimum, the potentials are written;

Lennard-Jones

$$U(r) = 4\epsilon \left[(\sigma/r)^{12} - (\sigma/r)^6 \right], \quad (1)$$

where σ is the value of r at $U(r) = 0$.

Exp: 6

$$U(r) = \frac{\epsilon}{1-6/\alpha} \left[\frac{6}{\alpha} e^{\alpha(1-r/r_m)} - \left(r_m/r \right)^6 \right], \quad (2)$$

where r_m is the value of r at the energy minimum and α a parameter which represents the steepness of the repulsive part of the function.

Kihara

$$U(r) = 4\epsilon \left[\left(\frac{\sigma - a}{r - a} \right)^{12} - \left(\frac{\sigma - a}{r - a} \right)^6 \right], \quad r > a \quad (3)$$

$$U(r) = \infty, \quad r \leq a.$$

Here the finite size of the molecule is taken into consideration by including a core diameter, a . (For the Lennard-Jones, $a = 0$.)

3. KINETIC THEORY EXPRESSIONS FOR THE TRANSPORT COEFFICIENTS

The kinetic theory for a dilute gas (see Appendix II) is formally complete [1], the Chapman-Enskog treatment of the Boltzmann equation giving the transport coefficients in terms of collision integrals which are functions of the gas dynamics and thus, of the intermolecular potential. It is the lack of knowledge of the latter which restricts the applicability of the kinetic theory expressions. The expressions used for the calculations are;

Viscosity (η)

$$\eta 10^7 = \frac{266.93 (MT)^{1/2}}{R^2 \Omega_{(2,2)}^* (T^*)} f_\eta \text{ g cm}^{-1} \text{ sec}^{-1}, \quad (4)$$

Thermal Conductivity (λ)

$$\lambda 10^7 = \frac{1989.1}{R^2} \frac{(T/M)^{1/2}}{\Omega^{(2,2)*}(T^*)} f_{\eta} \text{ cal cm}^{-1} \text{ sec}^{-1} \text{ deg}^{-1} \quad (5)$$
$$\equiv \frac{15}{4} \frac{\mathcal{R}}{M} \eta 10^7 \text{ cal cm}^{-1} \text{ sec}^{-1} \text{ deg}^{-1},$$

Self Diffusion (D)

$$D = \frac{0.002628}{P R^2} \frac{(T^3/M)^{1/2}}{\Omega^{(1,1)*}(T^*)} f_D \text{ cm}^2 \text{ sec}^{-1}, \quad (6)$$

where: M = molecular weight. (M = 39.948 for argon)

R = a distance parameter i.e., $R \equiv \sigma$ for the Lennard-Jones and Kihara, and $R \equiv r_m$ for the Exp: 6.

T = the absolute temperature, °K

\mathcal{R} = the gas constant per mole

$\Omega^{(2,2)*}(T^*)$ and $\Omega^{(1,1)*}(T^*)$ = the reduced collision integrals (reduced by dividing by the integrals for the rigid sphere case) at the reduced temperature T^* , where $T^* = T/(\epsilon/k)$ with k the Boltzmann constant.

f_{η} , f_{λ} , and f_D are terms accounting for higher mathematical approximations to η , λ , and D and are slowly varying functions of T^* which seldom differ from unity by more than about 0.5%. To be consistent with the accuracy of the experimental viscosity data at extreme temperatures, they can be omitted from Eqs. 4, 5, and 6 without significant error.

Tables of the collision integrals as a function of T^* for each of the potentials can be obtained from several sources (e.g., Refs. 1, 3), the numerical values of the integrals depend on the method of integration but it was verified that the choice of any particular set of tables made no significant difference to the results presented here. The tables used were, in fact, taken from Refs. 1, 3, and 11.

Note that Eq. 5 for the thermal conductivity is only applicable to monatomic gases. For polyatomic gases a separate correction is necessary for any internal degree of freedom. The coefficient of self-diffusion is an artificial quantity but is included as a further check on the potential and the force constants chosen.

4. METHOD OF CALCULATION AND RESULTS*

It was first necessary to determine the "best values" for the parameters for each of the functions. As has been stated, the preliminary calculations require the experimental viscosity coefficients only. The general method is now outlined.

For all functions, the first step was to substitute experimental values of η , corresponding temperatures T, and a sensible value for R into Eq. 4, hence obtaining $\Omega^{(2,2)*}(T)$ as a function of T. An interpolation routine next generated T^* by interpolating for these $\Omega^{(2,2)*}(T)$ values in the given set of $\Omega^{(2,2)*}(T^*)$ and T^* for the function in question. As $\epsilon/k = T/T^*$, ϵ/k was then found as a function of T for a particular R. Incidentally, these ϵ/k versus T curves were very sensitive to a change in the experimental viscosity coefficient and gave clear indications as to which sets of experimental data were internally consistent. (The viscosity coefficients at one atmosphere

* Much of this section is taken from the account in Ref. 13. The material is included for completeness.

were taken from Refs. 14-21; the same experimental data were used for all three potentials.) The procedure was repeated with several values of R , varying R by about 6% overall. It was found that the shape of the ϵ/k versus T^* curves indicated the best value of R . After selection of R and ϵ/k the potentials were compared by examining the deviation curves between theory and experiment.

The details are as follows:

4.1 LENNARD-JONES

The ϵ/k versus T curves for the σ 's of Table I obtained by following the above procedure are shown in Fig. 1. The theoretical collision integrals were taken from the tables in Ref. 3. The "best value" for σ was taken to be that associated with the curve with the least variation of ϵ/k over the widest temperature range allowing for the experimental scatter and interpolation error (about 1%). Trial and error selected $\sigma = 3.41 \text{ \AA}$. A corresponding best value of ϵ/k was about 125°K from Fig. 1 and this was adjusted to $\epsilon/k = 124.9^\circ\text{K}$ to obtain agreement between the experimental and calculated viscosity coefficients at 293°K. 293°K was chosen because the experimental viscosity coefficient at this temperature has been carefully measured by many workers [e.g., Refs. 16, 18, 20] with agreement of about 0.2%. Using the selected σ and ϵ/k , (Table II), a percentage deviation curve over the temperature range 90 to 1200°K was plotted, Fig. 2. It is to be noted in this figure that agreement between theory and experiment is substantially better than obtained by previous correlations [2-4].

TABLE II

Function	$\epsilon/k, ^\circ\text{K}$	$R, \text{\AA}$
L - J	124.9	3.41
Exp: 6	122.9	3.86
Kihara	139.8	3.35

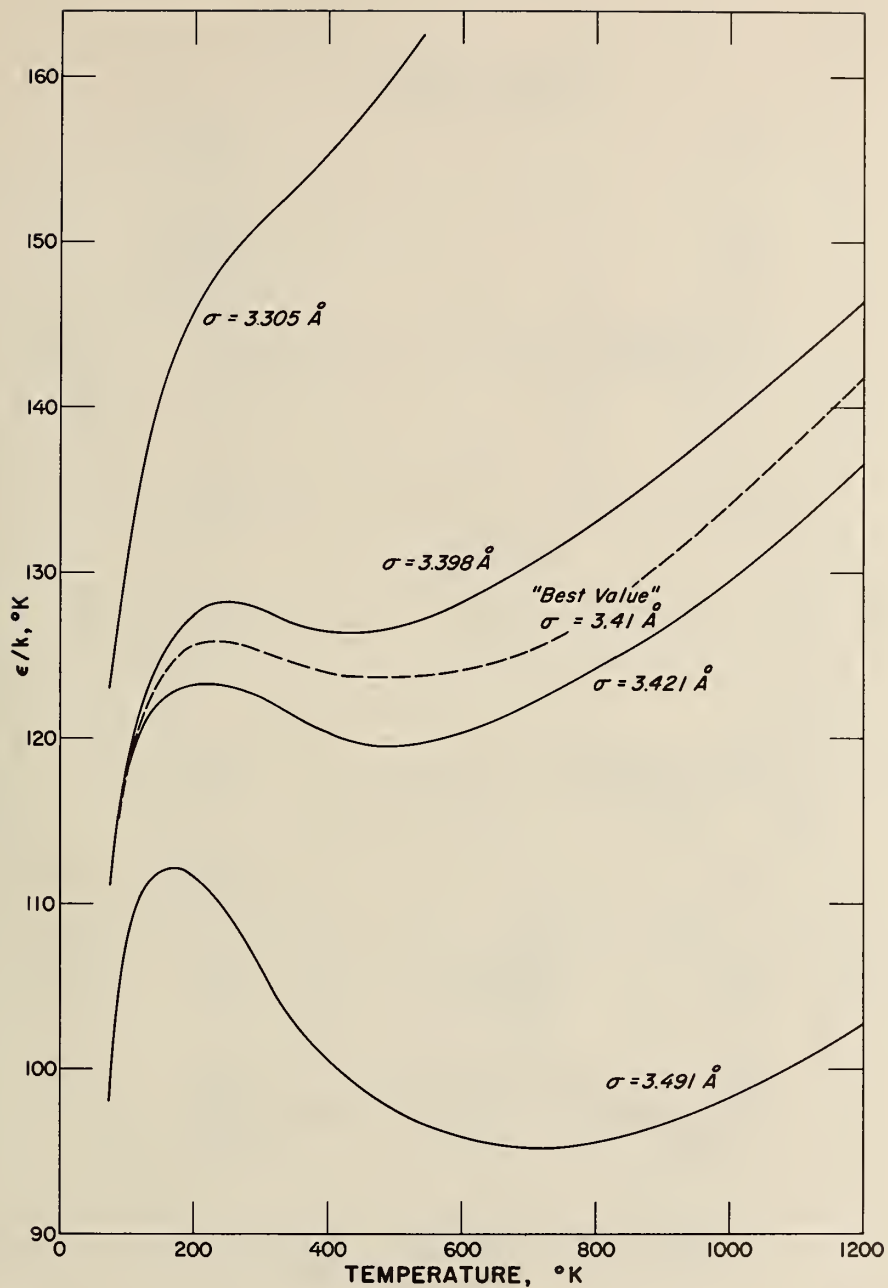
Table II. Best values of the parameters obtained by the method explained in the text.

4.2 EXP: 6

For this function α had to be selected first. It was initially taken as 12 and curves of ϵ/k versus T generated around $r_n = 3.5 \text{ \AA}$ using the collision integrals from Ref. 1, 11. When r_n was varied by about 6%, exactly the same pattern of curves were obtained as for the Lennard-Jones, Fig. 1. The curve for the best value of r_n for $\alpha = 12$ was selected as above and shown in Fig. 3. The procedure was repeated for $\alpha = 13, 14, 15, 16,$ and 17 with the same results, and corresponding curves for the best values are plotted in Fig. 3. It is seen that there is virtually no difference in the curves so $\alpha = 14$ was chosen to agree with Refs. 1, 11, and 12. Best values of ϵ/k and r_n were thus selected, (Table II) as outlined above and a deviation curve plotted, Fig. 2. The Exp: 6 function does not reproduce the data at high temperatures as well as the Lennard-Jones.

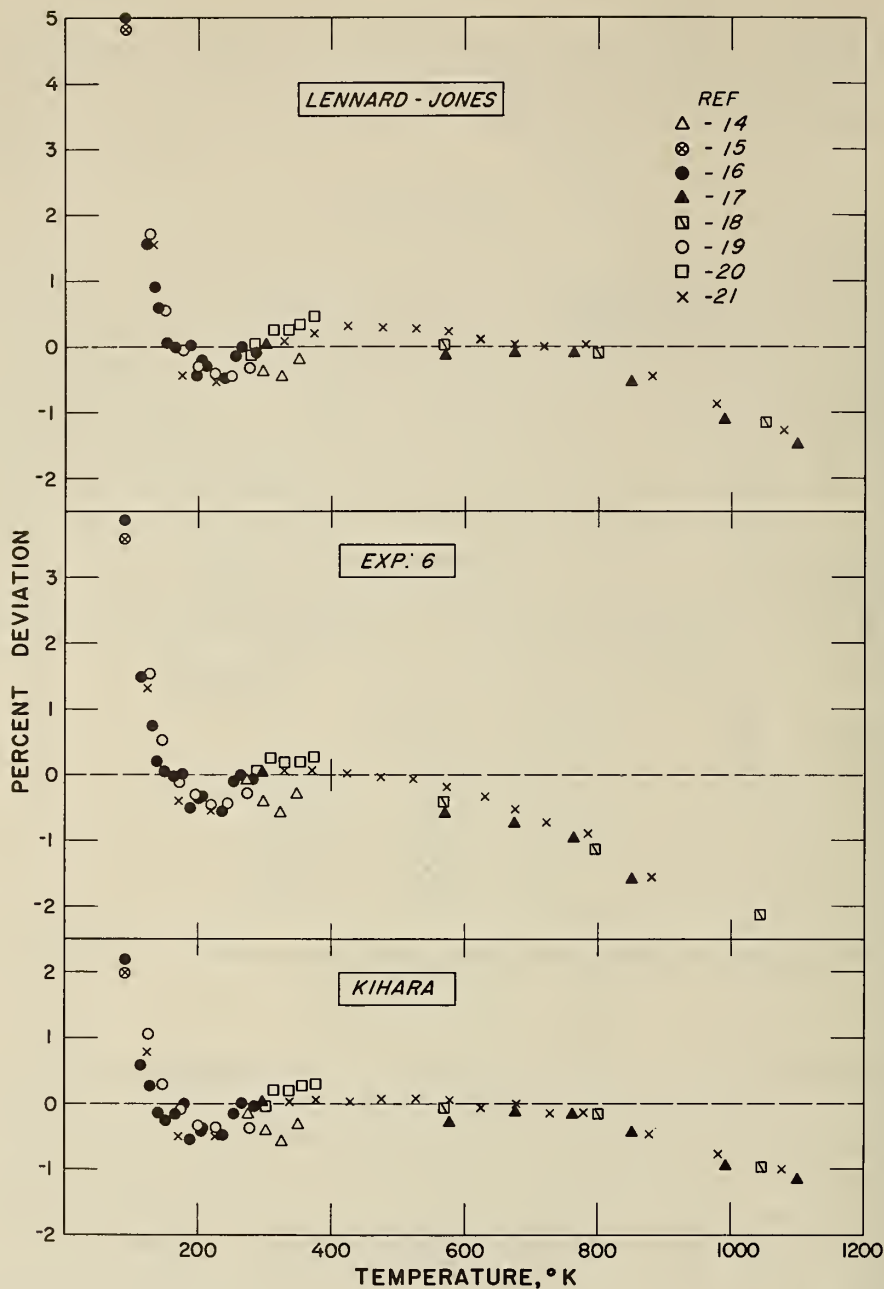
4.3 KIHARA

The variation of the reduced parameter γ (with γ defined as a/σ) was first investigated, letting γ be 0.1, 0.2, 0.3, 0.4, and 0.5 ($\gamma = 0$ is equivalent to the Lennard-Jones). The tables from Ref. 3 were used. It was found that the Lennard-Jones pattern of curves was only observed when $\gamma = 0.1$ and 0.2; the other values gave different patterns when σ was varied, ϵ/k was always high at the lower temperatures. Typical curves are shown in Fig. 4 and 5. γ was taken as 0.1 [3, 22] and a best value of ϵ/k found by fitting to 293°K as before, Table II. The deviation curve is shown in Fig. 2. It is seen that the Kihara appears to be slightly better than the Lennard-Jones overall and decidedly superior at low temperatures.



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Fig. 1. Variation of ϵ/k $^\circ K$ versus T $^\circ K$ for various best values of σ , selected by other authors, using the Lennard-Jones potential function. The dotted curve represents the best value of σ selected here. Similar curves were obtained for the Exp: 6 and Kihara functions when R was varied, except when $\gamma = 0.3, 0.4$ and 0.5 .



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Fig. 2. Percentage deviation plots $\left[\left(\frac{\eta_{\text{exp}} - \eta_{\text{calc}}}{\eta_{\text{calc}}} \right) \times 100 \right]$ of experimental and calculated viscosity coefficients for the three potential functions using the best values from Table II (Note: data from Ref. 14 is obtained by extrapolation to a pressure of one atmosphere).

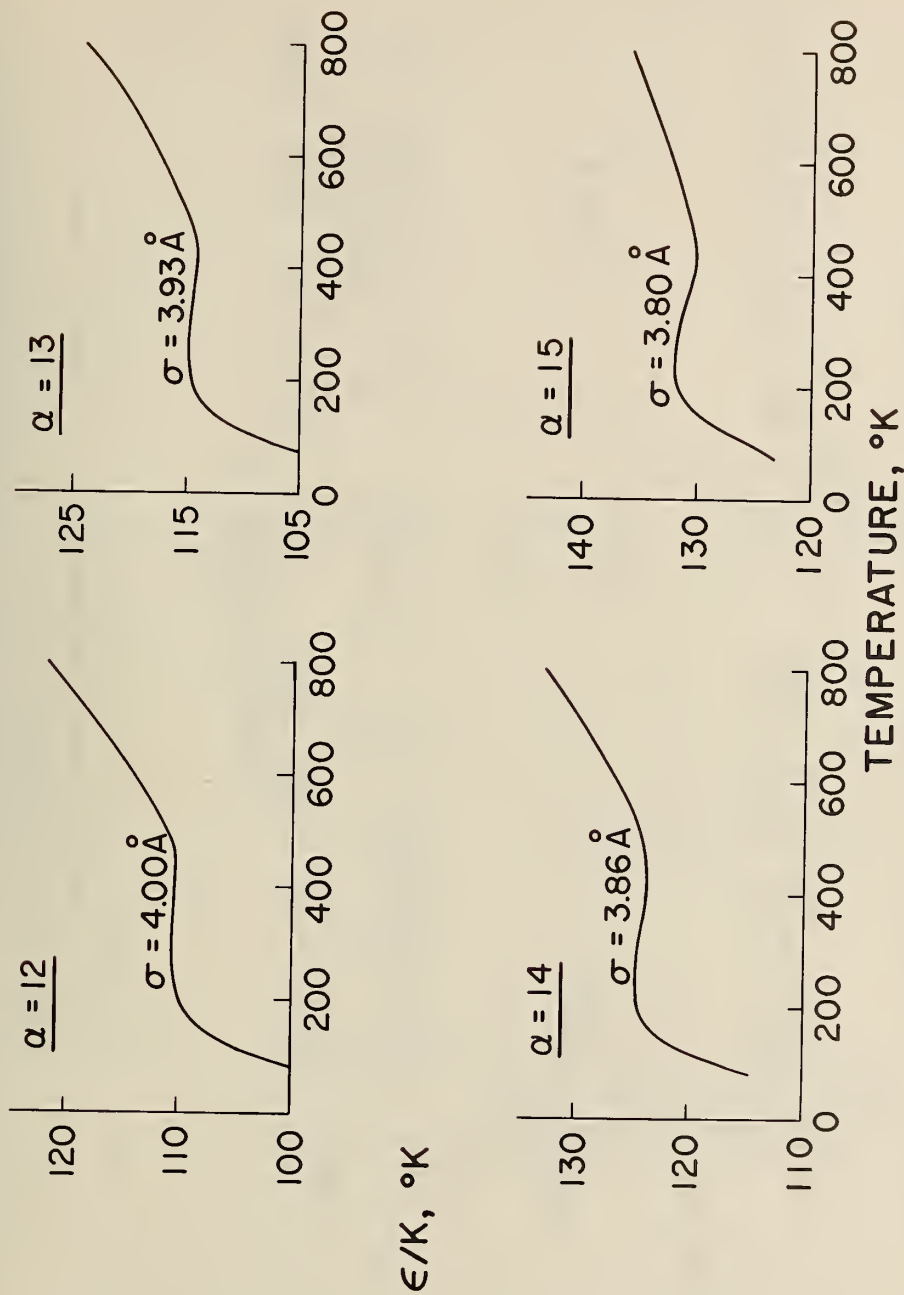


FIG. 3. Best values of r_n for various values of α - Exp: 6 potential. (Not shown: $\alpha = 16$ where $\sigma = 3.74 \text{ \AA}$, and $\alpha = 17$ where $\sigma = 3.67 \text{ \AA}$.)

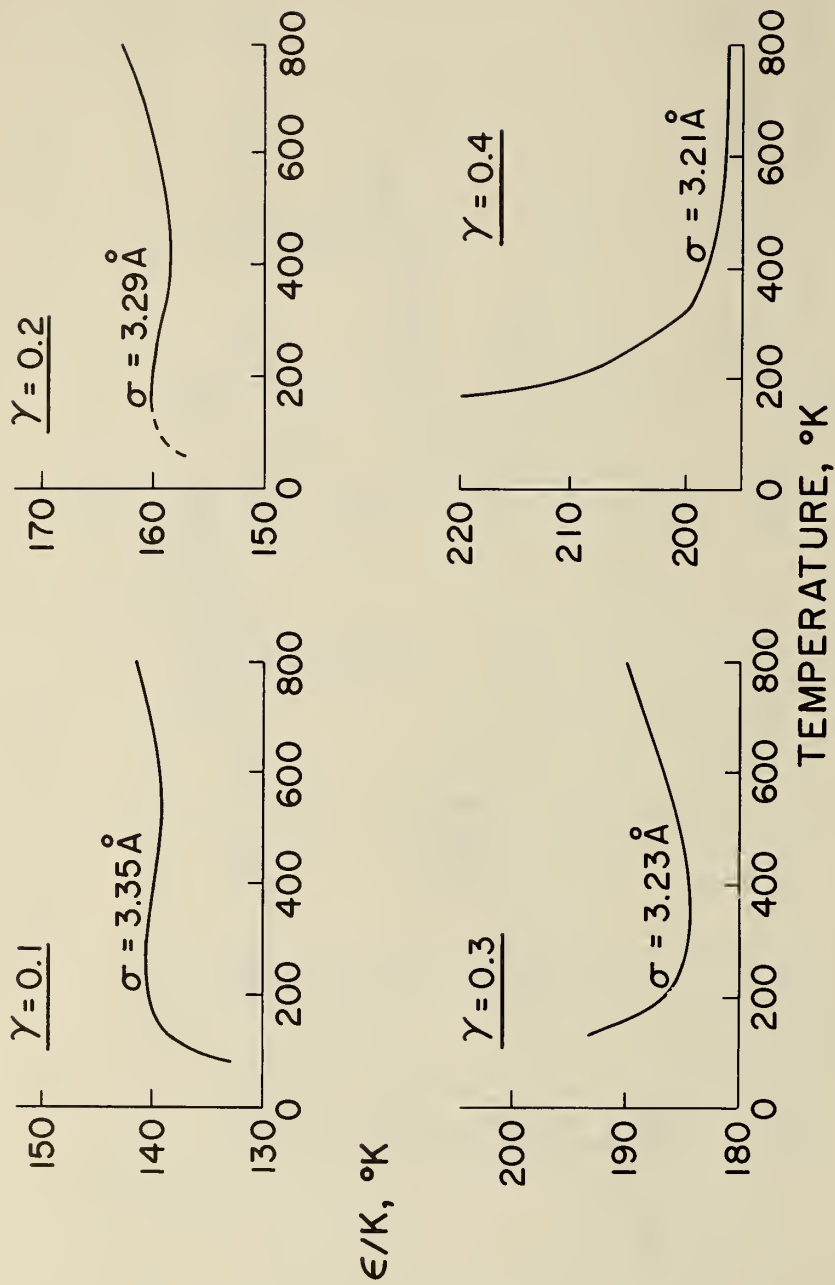


Fig. 4. Best values of σ for $\gamma = 0.1, 0.2, 0.3,$ and 0.4 - Kihara potential. A choice of a best value becomes less meaningful if γ is fixed at a value greater than 0.2 . (See Fig. 5.)

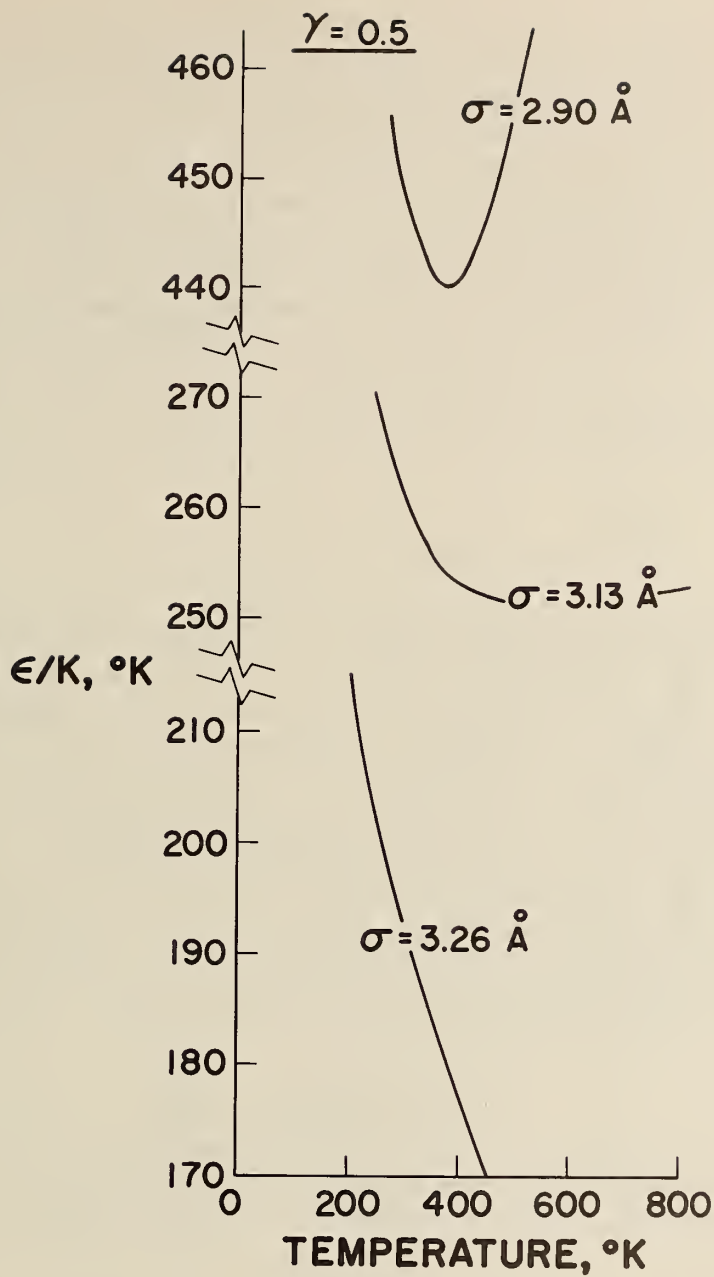


Fig. 5. The pattern of curves for $\gamma = 0.5$ is illustrated by $\sigma = 2.90 \text{ \AA}$, $\sigma = 3.13 \text{ \AA}$ and $\sigma = 3.26 \text{ \AA}$.

The conclusion from the above is that care must be taken to rationalize the choice of R and ϵ/k when comparing one potential function with another. It is seen that less than a 6% variation in R will make comparisons over a wide temperature range meaningless, a factor neglected by many authors. It is shown here that for argon the Lennard-Jones gives a much better correlation with experiment than before realized and that the Kihara potential is even better. In fact, in the temperature range 100 to 1500°K, the correlation is almost as good as could be obtained from an empirical polynomial fit [23]. The deviations at high and low temperatures have not been satisfactorily explained. It has been suggested that the experimental viscosity coefficients are too low at the higher temperatures [5, 11]. It will be shown that correlation curves for the thermal conductivity and self-diffusion coefficients tend to support this. The deviations at the lower temperatures are not likely to be caused by quantum effects and the reason is still really unsolved (see, however, Ref. 5). Although the Kihara is thus recommended as the most realistic function it is clear that, in practice, there is really very little to choose from among these functions. In fact, the method illustrates the arbitrariness in choosing a potential function.

5. THERMAL CONDUCTIVITY

The experimental and kinetic theory viscosity coefficients have been shown to be satisfactorily correlated with the Kihara potential and the parameters given in Table II. The corresponding thermal conductivity expression, Eq. 5 should be equally applicable for dilute argon. This was tested by plotting the deviation curve of the resulting theoretical coefficient from the experimental coefficient, Fig. 6 (the experimental values, at one atmosphere, were taken from Refs. 24-32). As is well known the scatter in these experimental values is large, especially at temperatures outside the range 200 to 600°K. It is seen, however, that the kinetic theory, with the Kihara potential gives an adequate correlation. Of course, as Eq. 5 was used, the same deviations observed with the viscosity correlation will be introduced into the thermal conductivity correlation. This is apparent at low temperatures, but not at high temperatures. The experimental viscosity coefficients at these high temperatures may be low, but thermal conductivity data are not reliable enough to give any definite conclusions. Thus, tables of the thermal conductivity coefficients computed with the Kihara are reasonable above 100°K, although there must be uncertainty above 1200°K.

6. SELF-DIFFUSION

For completeness the self-diffusion coefficient, Eq. 6, was included and a deviation curve plotted, Fig. 7. The experimental data, taken from Refs. 33-37, are scarce and so the correlation, although reasonably satisfactory, is not very meaningful.

7. CONCLUSION

The above three deviation plots (Figs. 2, 6, and 7) indicate that the kinetic theory expressions for the transport coefficients can be used to compute tables for dilute argon, at least to within about four percent between 100 and 1200°K. Thus, tables of the viscosity and thermal conductivity coefficients for temperatures from 100 to 2000°K have been computed and are given in Appendix I. It is recognized that the error in any fitted function tends to increase as the independent variable, temperature in this case, increases beyond the range of the data.

8. ACKNOWLEDGEMENT

This work was supported by a grant from NASA. Thanks are given to Gregg Childs for assistance with the preparation of the data, and to Max Klein, J. G. Hust, R. B. Stewart, and R. D. McCarty for valuable discussions.

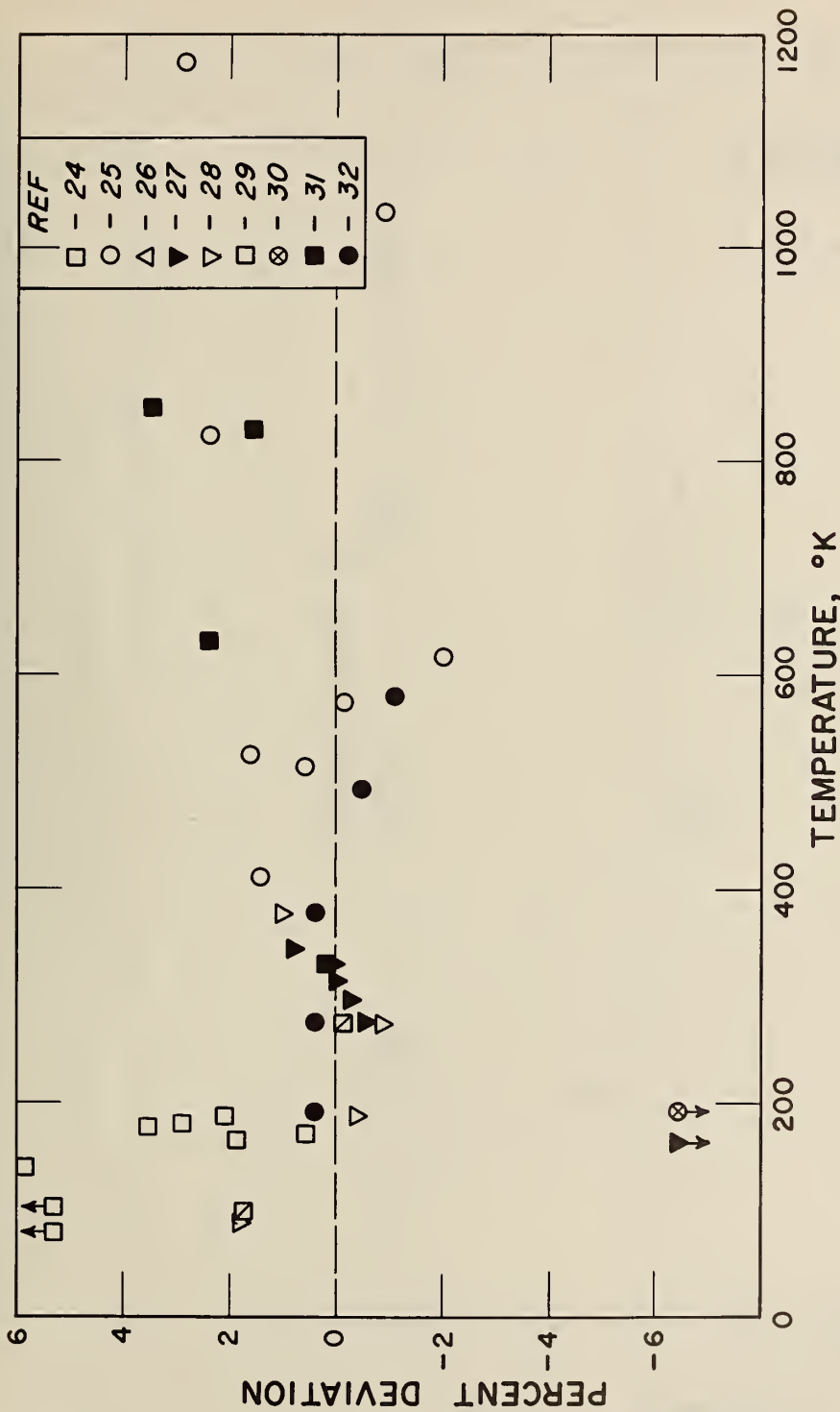


Fig. 6. Deviation plot $\left[\left(\frac{\lambda_{exp} - \lambda_{calc}}{\lambda_{calc}} \right) \times 100 \right]$ for the experimental thermal conductivity coefficients and those computed from Eq. 5 with the Kihara potential collision integrals from Ref. 3. Note that the trend at temperatures greater than 700°K found with the viscosity coefficients is not repeated.

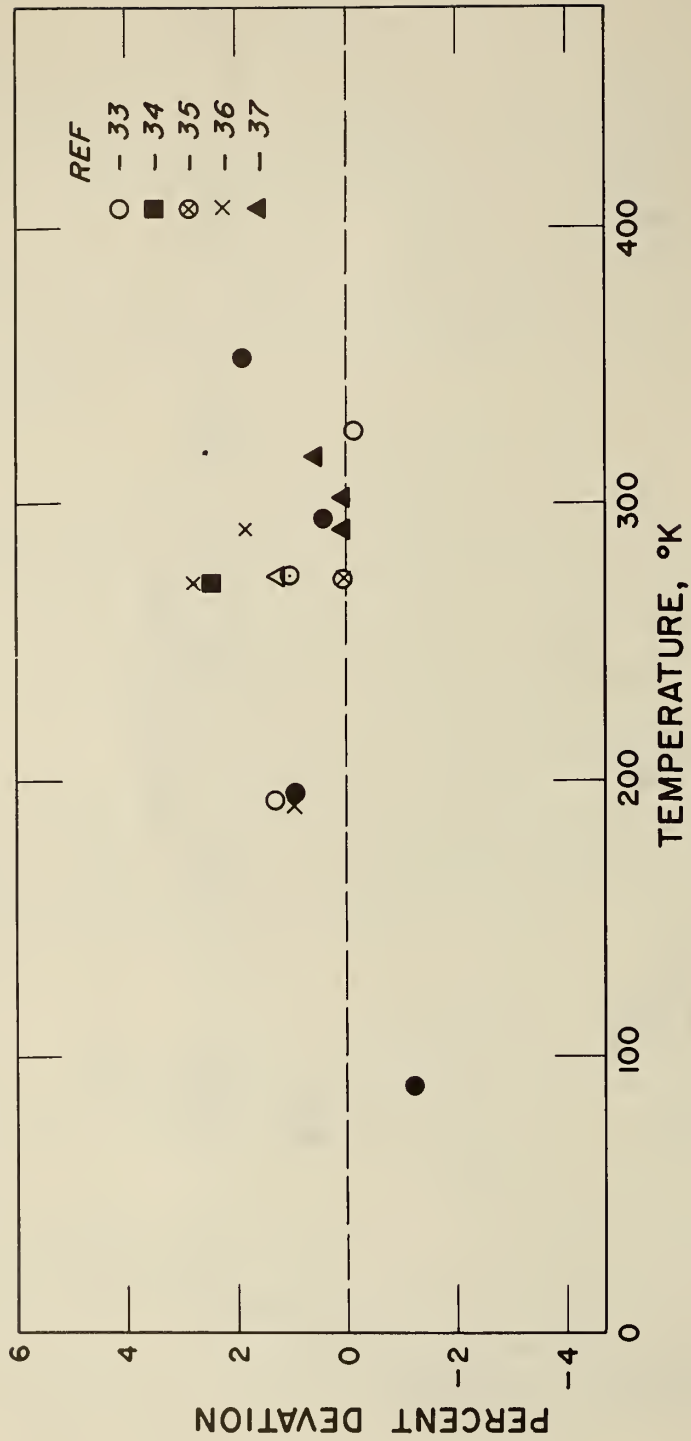


Fig. 7. Deviation plot $\left[\left(\frac{D_{exp} - D_{calc}}{D_{calc}} \right) \times 100 \right]$ for the experimental diffusion coefficients computed on the same basis as the viscosity and thermal conductivity coefficients.

9. APPENDIX I

THE VISCOSITY AND THERMAL CONDUCTIVITY COEFFICIENTS OF DILUTE ARGON FROM 100 TO 2000°K

The tables were computed from Eqs. 4 and 5 using the Kihara collision integrals from Ref. 3 and letting $\gamma = 0.1$, $\sigma = 3.35\text{\AA}$, and $\epsilon/k = 139.8^\circ\text{K}$.

VISCOSITY AND THERMAL CONDUCTIVITY OF GASEOUS ARGON*

TEMPERATURE	VISCOSITY	THERMAL CONDUCTIVITY	TEMPERATURE	VISCOSITY	THERMAL CONDUCTIVITY
K	G/CM-SEC $\eta \times 10^6$	CAL/CM-SEC-DEG $\lambda \times 10^6$	K	G/CM-SEC $\eta \times 10^6$	CAL/CM-SEC-DEG $\lambda \times 10^6$
			500	334.9	62.5
			510	339.8	63.4
			520	344.5	64.3
			530	349.2	65.1
			540	353.8	66.0
			550	358.3	66.8
			560	362.9	67.7
			570	367.4	68.5
			580	371.9	69.4
			590	376.4	70.2
			600	380.8	71.0
			610	385.1	71.8
			620	389.5	72.6
			630	393.8	73.4
			640	398.0	74.2
			650	402.3	75.0
			660	406.5	75.8
			670	410.6	76.6
			680	414.8	77.4
			690	418.9	78.1
			700	423.0	78.9
			710	427.1	79.7
			720	431.1	80.4
			730	435.1	81.2
			740	439.1	81.9
			750	443.1	82.6
			760	447.1	83.4
			770	451.0	84.1
			780	454.7	84.8
			790	458.4	85.5
			800	462.5	86.3
			810	466.3	87.0
			820	470.0	87.7
			830	473.8	88.4
			840	477.6	89.1
			850	481.3	89.8
			860	485.0	90.5
			870	488.7	91.2
			880	492.4	91.9
			890	496.1	92.5
			900	499.8	93.2
			910	503.5	93.9
			920	507.1	94.6
			930	510.7	95.3
			940	514.4	95.9
			950	518.0	96.6
			960	521.5	97.3
			970	525.1	97.9
			980	528.7	98.6
			990	532.2	99.3
100	82.5	15.4			
110	90.6	16.9			
120	98.7	18.4			
130	106.8	19.9			
140	114.8	21.4			
150	122.7	22.9			
160	130.4	24.3			
170	138.1	25.8			
180	145.7	27.2			
190	153.1	28.6			
200	160.4	29.9			
210	167.7	31.3			
220	174.6	32.6			
230	181.6	33.9			
240	188.5	35.2			
250	195.2	36.4			
260	201.8	37.6			
270	208.3	38.9			
280	214.7	40.0			
290	220.9	41.2			
300	227.1	42.4			
310	233.3	43.5			
320	239.4	44.6			
330	245.2	45.7			
340	251.0	46.8			
350	256.7	47.9			
360	262.4	48.9			
370	268.1	50.0			
380	273.6	51.0			
390	279.1	52.1			
400	284.5	53.1			
410	289.8	54.1			
420	295.1	55.0			
430	300.4	56.0			
440	305.5	57.0			
450	310.5	57.9			
460	315.5	58.9			
470	320.5	59.8			
480	325.4	60.7			
490	330.3	61.6			

* Calculated for the dilute gas by the Kihara potential, with $\gamma = .1$, $\sigma = 3.35 \text{ \AA}$, $\epsilon/k = 139.8^\circ\text{K}$.
Figure 11 defines temperature and pressure range for the dilute gas.

VISCOSITY AND THERMAL CONDUCTIVITY OF GASEOUS ARGON

TEMPERATURE	VISCOSITY	THERMAL CONDUCTIVITY	TEMPERATURE	VISCOSITY	THERMAL CONDUCTIVITY
K	G/CM-SEC	CAL/CM-SEC-DEG	K	G/CM-SEC	CAL/CM-SEC-DEG
	$\eta \times 10^6$	$\lambda \times 10^6$		$\eta \times 10^6$	$\lambda \times 10^6$
1000	535.6	99.9	1500*	695.4	129.7
1010	539.0	100.5	1510*	698.4	130.3
1020	542.6	101.2	1520*	701.3	130.8
1030	546.0	101.8	1530*	704.3	131.4
1040	549.4	102.5	1540*	707.2	131.9
1050	552.9	103.1	1550*	710.1	132.4
1060	556.3	103.8	1560*	713.0	133.0
1070	559.7	104.4	1570*	715.9	133.5
1080	563.0	105.0	1580*	718.9	134.1
1090	566.4	105.6	1590*	721.7	134.6
1100	569.8	106.3	1600*	724.6	135.2
1110	573.1	106.9	1610*	727.5	135.7
1120	576.5	107.5	1620*	730.4	136.2
1130	579.8	108.1	1630*	733.3	136.8
1140	583.1	108.8	1640*	736.1	137.3
1150	586.4	109.4	1650*	739.0	137.8
1160	589.7	110.0	1660*	741.8	138.4
1170	593.0	110.6	1670*	744.7	138.9
1180	596.3	111.2	1680*	747.4	139.4
1190	599.5	111.8	1690*	750.2	139.9
1200*	602.8	112.4	1700*	753.0	140.4
1210*	606.0	113.0	1710*	755.8	141.0
1220*	609.2	113.6	1720*	758.6	141.5
1230*	612.4	114.2	1730*	761.6	142.1
1240*	615.5	114.8	1740*	764.4	142.6
1250*	618.8	115.4	1750*	767.2	143.1
1260*	622.0	116.0	1760*	770.0	143.6
1270*	625.2	116.6	1770*	772.7	144.1
1280*	628.3	117.2	1780*	775.5	144.6
1290*	631.4	117.8	1790*	778.3	145.2
1300*	634.6	118.4	1800*	781.0	145.7
1310*	637.7	118.9	1810*	783.8	146.2
1320*	640.8	119.5	1820*	786.5	146.7
1330*	643.9	120.1	1830*	789.3	147.2
1340*	647.0	120.7	1840*	792.0	147.7
1350*	650.1	121.3	1850*	794.7	148.2
1360*	653.2	121.8	1860*	797.4	148.7
1370*	656.3	122.4	1870*	800.1	149.2
1380*	659.4	123.0	1880*	802.8	149.7
1390*	662.4	123.6	1890*	805.5	150.2
1400*	665.5	124.1	1900*	808.2	150.7
1410*	668.5	124.7	1910*	810.8	151.2
1420*	671.6	125.3	1920*	813.5	151.7
1430*	674.6	125.8	1930*	816.2	152.2
1440*	677.6	126.4	1940*	818.8	152.7
1450*	680.5	126.9	1950*	821.5	153.2
1460*	683.5	127.5	1960*	824.1	153.7
1470*	686.4	128.0	1970*	827.0	154.3
1480*	689.5	128.6	1980*	829.7	154.7
1490*	692.5	129.2	1990*	832.3	155.2
			2000*	834.9	155.7

* There is some uncertainty in these values, see text.

10. APPENDIX II

AN ESTIMATE OF THE EXTENT OF VALIDITY OF THE TRANSPORT PROPERTY TABLES

It is interesting to estimate the experimental conditions for which the tables are valid. These tables are, of course, only applicable when the gas is dilute, where a dilute gas can be defined as a gas in which only binary molecular collisions are of consequence: in other words, the gas dynamics and the intermolecular interactions are completely determined by the collisions and interactions between pairs of molecules alone. The gas must not be so rarefied, however, that gas-wall collisions are of any account. Furthermore, the viscosity coefficient, Eq. 4, must be independent of pressure, Thus, the tables are inapplicable for experimental conditions where Eq. 4 no longer reproduces the experimental data (to within about 0.5% in the temperature range of 100 to 1500°K). Also if Eq. 4 does not reproduce the data the gas cannot be regarded as dilute as defined. Therefore, a straightforward procedure could be to compare experimental viscosity coefficients at pressures greater than one atmosphere with the equivalent theoretical values from Eq. 4 at the corresponding temperatures, and observe when the percentage deviation becomes appreciable. This was indeed done with the experimental data from Refs. 1, 14, 38-40 and percentage deviations calculated for the given temperatures and pressures. A qualitative, but simple, way to examine the results was to plot the percentage deviations versus a reduced mean free path λ_R . As the results were only qualitative, the elementary kinetic theory expression for the mean free path was used:

$$\lambda = \frac{3 \eta \mathcal{R}}{P \bar{v} M} \quad \text{cm} , \quad (7)$$

where \mathcal{R} is the gas constant per mole. P the pressure in dynes and $\bar{v} = (8RT/\pi M)^{1/2}$. λ was reduced by dividing by the mean free path of argon at one atmosphere and at 293°K, ($\eta_{293^\circ} = 2.23 \times 10^{-4} \text{ g cm}^{-1} \text{ sec}^{-1}$). We then obtain:

$$\lambda_R = 257 \eta (T)^{1/2} / P , \quad (8)$$

where η is the viscosity coefficient at temperature T and pressure P (P now in atmospheres). Figure 8 shows the plot and it is seen that there is an appreciable rise at $\lambda_R = 0.07$ and percentage deviation = 1%. It is reasonable to take this point as the extent of validity of Eq. 4 and as the limit of dilution. The limiting value of λ was then used (Eq. 7) to construct Fig. 9 which shows an approximate limiting P - T curve above which argon cannot be considered dilute. A curve for a 2% deviation ($\lambda_R = 0.05$) is also included. Unfortunately, the experimental work was not carried out at high temperatures and the higher temperature data available (Ref. 38) do not agree too well with other workers, Fig. 8.

The above exercise is very qualitative; even for argon there is uncertainty in the values of force-constants, the experimental data are not as extensive as desired and obviously Eq. 7 is only approximate. Also, more seriously, one cannot be sure how the failure of the potential function to exactly reproduce the data, especially at the higher temperatures, is contributing to the rise in the curve in Fig. 8. However, Fig. 9 is certainly reasonable and it is expected that similar curves would approximately represent the limit of dilution for simple gases at temperatures and pressures removed from the critical and boiling regions.

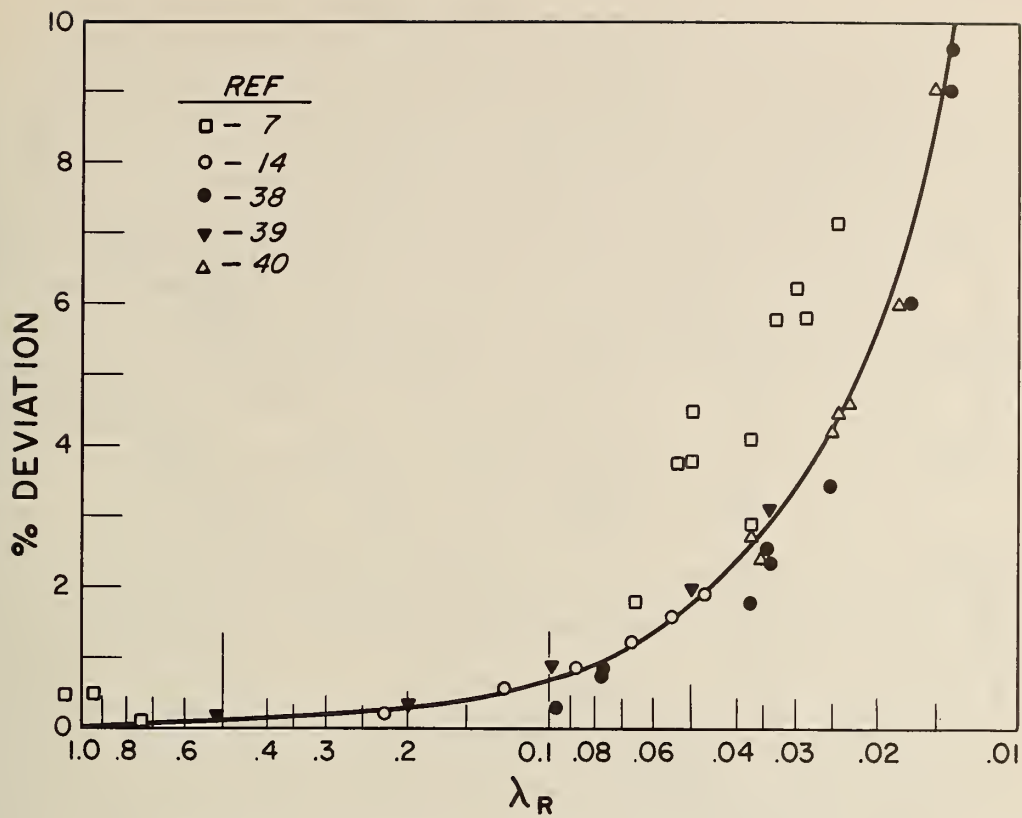


Fig. 8. Plot of the percentage deviation of the experimental and theoretical viscosity coefficients calculated from Eq. 4 versus the reduced mean free path λ_R , Eq. 8.

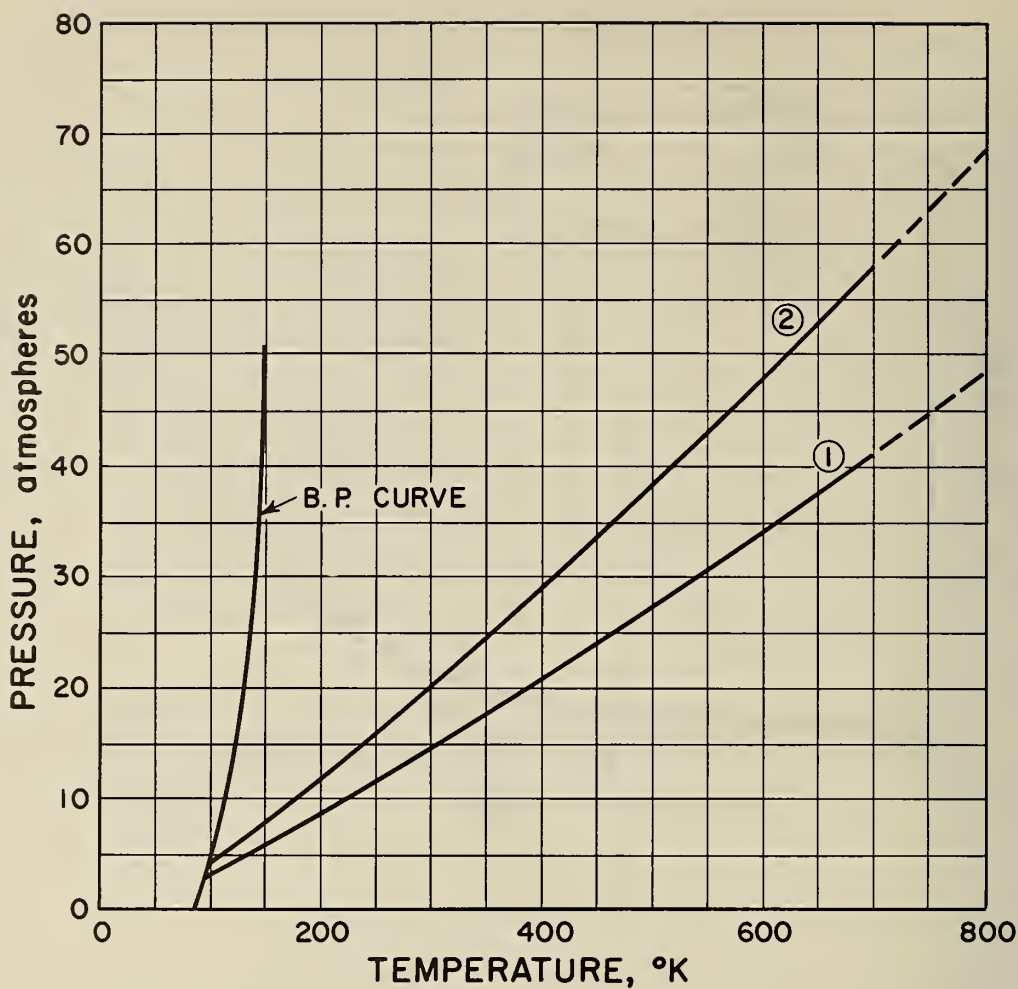


Fig.9. Plot of the pressure versus temperature limiting dilution curves for argon. The gas can be assumed to be dilute for all pressures and temperatures below curve 1. Curve 1 was constructed by comparing experimental viscosity coefficients with the theoretical, Eq. 4, and is the curve for a 1% deviation (See explanation in the text). Curve 2 which represents a 2% deviation between theory and experiment is included for comparison. Curve 1 represents a constant density curve of 0.025 g cm^{-3} .

THE SELECTION OF FORCE CONSTANTS

Force constants, for a particular potential function, are often determined from as many different experimental sources as possible [1, 2, 4], but it is well known that constants obtained in this general way often lead to anomalous results when they are applied to a single experiment. For example, values of the constants found from, say, equation of state, Joule-Thomson, or crystal structure data, when applied to the transport coefficients do not always give as good results as the values of the constants determined from transport data alone. This is to be expected because the potential function is not exact. The inconsistency has not been discussed quantitatively (see, however, Ref. 41).

The above can be illustrated by comparing the curves for the variation of ϵ/k with T for the Lennard-Jones at several values of σ (Fig. 1) with equivalent curves calculated from second virial coefficients, $B(T)$,

$$B(T) = b_0 B^*(T^*) \quad (9)$$

$$b_0 = \frac{2}{3} \pi N \sigma^3$$

where N = Avagadro's number and $B^*(T^*)$ is the reduced second virial at the reduced temperature T^* . (In Ref. 22, ϵ/k and b_0 were found to depend on the temperature but the temperature range was restricted from 90 to 400°K.)

It was straightforward to find the variation of ϵ/k with T at constant σ by following the viscosity procedure and referring to the reduced second virials in Ref. 1. The experimental data are given in Ref. 22. The resulting curves are shown in Fig. 10. When compared to Fig. 1 over a similar temperature range it is seen that the pattern is different and that the second virial does not appear to be as sensitive to a change in σ as the viscosity coefficient. A "best value" for ϵ/k in the 90 to 400°K temperature range is approximately $\epsilon/k = 119^\circ\text{K}$. Thus, according to Figs. 1 and 10 it would be best not to interchange the values of the parameters. However, it is instructive to see if a less sensitive relation exists between η and $B(T)$. η and $B(T)$ can be related if a common ϵ/k exists because from Eqs. 4 and 9,

$$1/\Omega^{(2,2)*}(T^*) = F(B^*(T^*)) \quad (10)$$

From the experimental data one can calculate,

$$1/\Omega^{(2,2)*}(T^*) = \eta_{\text{exp}} \frac{10^7 \sigma^2}{266.93} (MT)^{1/2} \quad (11)$$

and

$$B^*(T^*) = B_{\text{exp}} / \frac{2}{3} \pi N \sigma^3 \quad (12)$$

This has been done for three values of σ and the results drawn as points in Fig. 11. The full curve in Fig. 11 represents Eq. 10 constructed from the theoretical tables in Refs. 1 and 3. The sources of the viscosity coefficient experimental data are given in Section 4 and the second virial in Ref. 22. It is seen that when $\sigma = 3.42\text{\AA}$, agreement between theory and experiment is excellent and thus it would appear that a consistent pair of values for ϵ/k and σ would result from independent examination of both viscosity and second virial experimental data, but clearly, this latter method is not as sensitive as the former.

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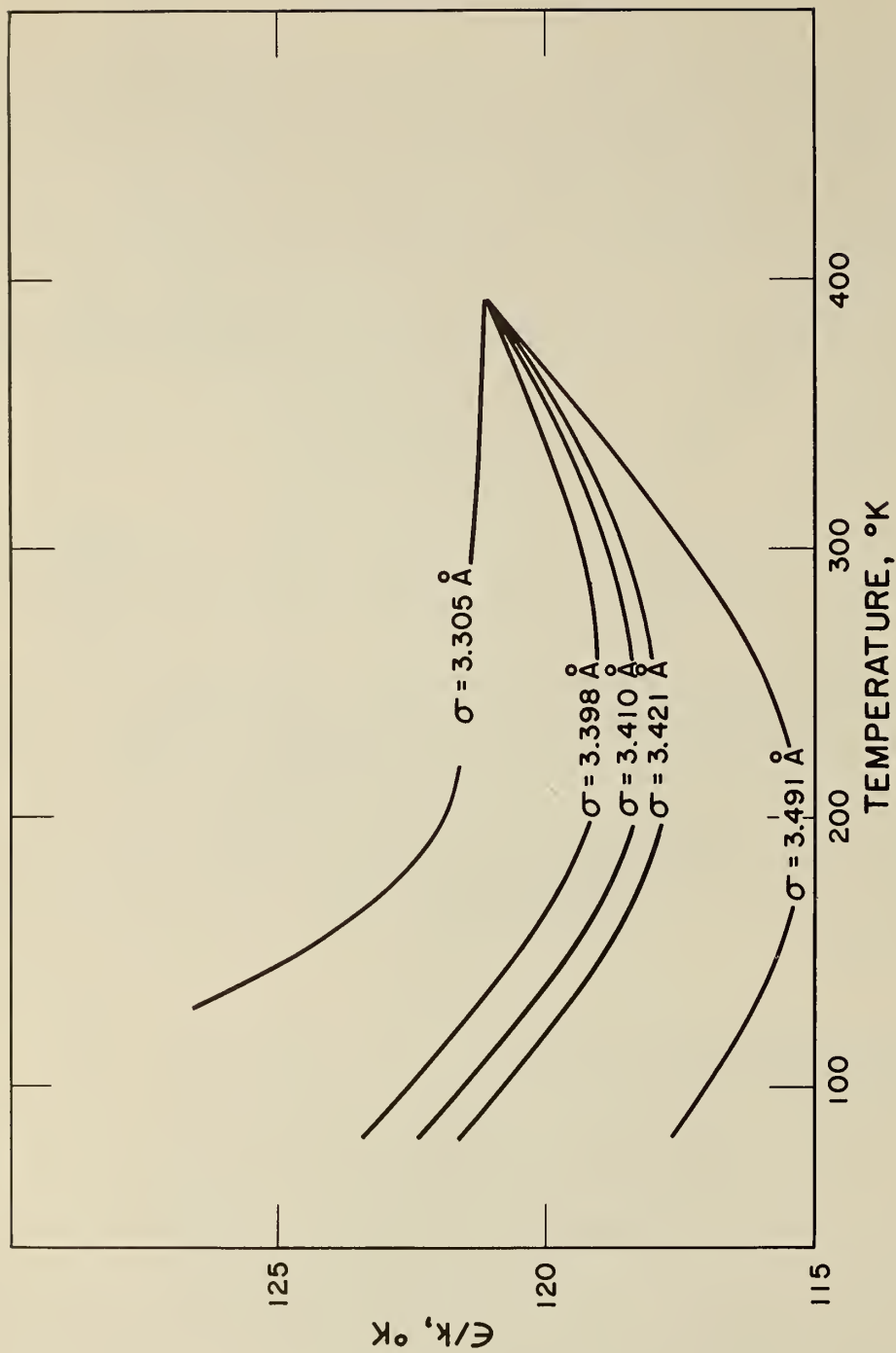


Fig. 10. Plot of the variation of ϵ/k versus temperature for several values of σ , (Table II). The variation was found with the second virial coefficients taken from Ref. 22, and Eq. 9, following the procedure outlined in Section 4.

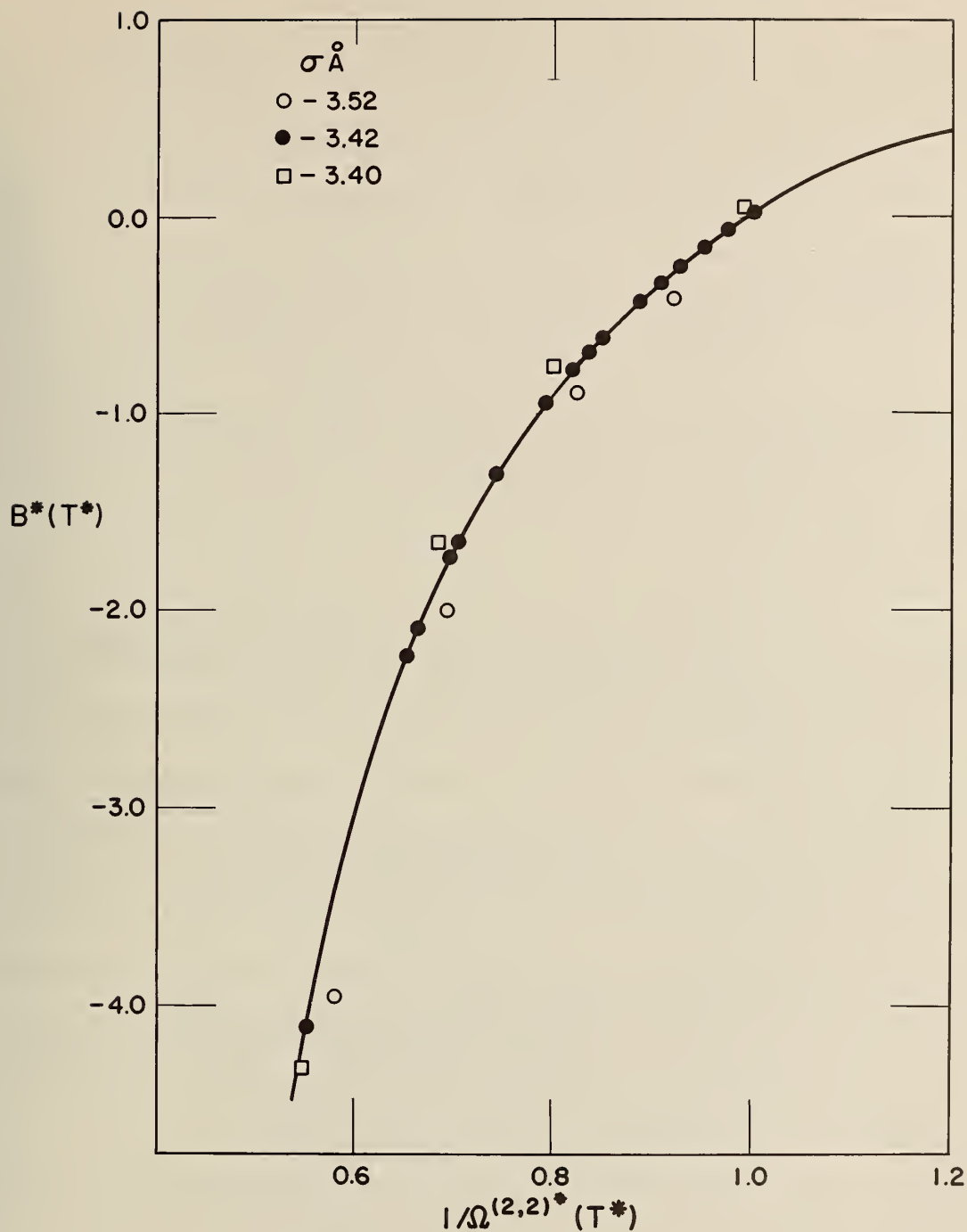


Fig. 11. Plot of $B^*(T^*)$ versus $1/\Omega^{(2,2)*}(T^*)$. The solid curve was generated from Eq. 10 using the second virial and collision integral tables from Refs. 1 and 3. The points represent the curve from experimental data with three values of σ .

12. REFERENCES

1. J. O. Hirschfelder, C. F. Curtiss, and R. B. Bird, *MOLECULAR THEORY OF GASES AND LIQUIDS*, John Wiley & Sons, New York (1964) Second Printing.
2. R. A. Svehla, Estimated Viscosities and Thermal Conductivities of Gases at High Temperatures, Natl. Aeronaut. Space Admin. Tech. Rept. R-132 (1962), DDC AD 272 963.
3. J. A. Barker, W. Fock, and F. Smith, Calculation of Gas Transport Properties and the Interaction of Argon Atoms, *Phys. Fluids* 7, 897-903 (1964).
4. J. Hilsenrath, C. W. Beckett, W. S. Benedict, L. Fano, H. J. Hoge, J. F. Masi, R. L. Nuttall, Y. S. Touloukian, and H. W. Woolley, Tables of Thermal Properties of Gases, Natl. Bur. Standards Circ. No. 564 (Nov 1955).
5. R. J. Munn and F. J. Smith, Interaction of the Inert Gases II, Univ. Maryland, Inst. Molecular Phys. IMP-NASA-47 (May 1965)
6. J. H. Milligan and P. E. Liley, Lennard-Jones Potential Parameter Variation as Determined from the Viscosity Data for Twelve Gases, *Am. Soc. Mech. Engrs. Paper No. 64-HT-20* (1964).
7. J. Kestin and J. H. Whitelaw, A Relative Determination of the Viscosity of Several Gases by the Oscillating Disk Method, *Physica* 29, 335-56 (1963).
8. E. Whalley and W. G. Schneider, The Lennard-Jones 12-6 Potential and the Viscosity of Gases, *J. Chem. Phys.* 20, 657-61 (1952).
9. J. S. Rowlinson, A Test of Kihara's Intermolecular Potential, *Mol. Phys.* 9, No. 2, 197-98 (1965).
10. J. P. O'Connell and J. M. Prausnitz, Applications of the Kihara Potential to Thermodynamic and Transport Properties of Gases, *ADVANCES IN THERMOPHYSICAL PROPERTIES AT EXTREME TEMPERATURES AND PRESSURES*, 19-31 *Am. Soc. Mech. Engr.*, New York (1965).
11. E. A. Mason and W. E. Rice, The Intermolecular Potentials for Some Simple Non-polar Molecules, *J. Chem. Phys.* 22, 843-51 (1954).
12. A. G. de Rocco and J. O. Halford, Intermolecular Potentials of Argon, Methane and Ethane, *J. Chem. Phys.* 28, 1152-54 (1958).
13. H. J. M. Hanley, Comparison of the Lennard-Jones, Exp: 6, and Kihara Potential Functions Using Viscosity Data of Dilute Argon, to be published in *J. Chem. Phys.* 1966.
14. A. Michels, A. Botzen, and W. Schuurman, The Viscosity of Argon at Pressures up to 2000 Atmospheres, *Physica* 20, 1141-48 (1954).
15. A. Van Itterbeek and O. Van Paemel, Measurements on the Viscosity of Argon Gas at Room Temperature and Between 90° and 55°K, *Physica* 5, 1009-11 (1938).
16. H. L. Johnston and E. R. Grilly, Viscosities of Carbon Monoxide, Helium, Neon, and Argon Between 80° and 300°K Coefficients of Viscosity, *J. Phys. Chem.* 46, 948-63 (1942).
17. M. Trautz and R. Zink, Die Reibung, Wärmeleitung und Diffusion in Gasmischungen. XII. Gasreibung bei höheren Temperaturen, (The Viscosity, Heat Conduction and Diffusion of Gas Mixtures. XII. The Viscosity of Gases at High Temperatures), *Ann. Physik* 7, 427-52 (1930).
18. V. Vasilescu, Recherches Experimentales sur la Viscosite des Gaz aux Temperatures Elevees, (Experimental Research on the Viscosity of Gas at High Temperatures), *Ann. Phys. (Paris)* 20, 137-76 (1945).
19. G. P. Filippova and I. P. Ishkin, The Viscosity of Air, Nitrogen, and Argon at Low Temperatures and at Pressures up to 150 Atmospheres, *Inzh. Fiz. Zh. Akad. Nauk Belorussk SSR* 4, 105-09 (1961).
20. R. Wobser and F. Müller, Die innere Reibung von Gases und Dämpfen und ihre Messung im Höppler-Viskosimeter, (The Viscosity of Gases and Vapors and their Measurement with the Hoppler Viscometer), *Kolloid-Beih.* 52, 165-276 (1941).
21. F. G. Keyes, The Heat Conductivity, Viscosity, Specific Heat and Prandtl Numbers for Thirteen Gases, *Mass. Inst. of Technol., Proj. Squid, Tech. Rept. No. 37* (1952), DDC ATI 167 173.
22. A. L. Gosman, Thermodynamic Properties of Argon in the Liquid and Gaseous State for Temperatures from the Triple Point to 300°K and Pressures to 1000 Atmospheres, *State Univ. of Iowa, Iowa City, Ph. D. Thesis* (1965).

23. Thermophysical Properties Research Center, DATA BOOK, Vol. 2, Purdue University, Lafayette, Ind. (1964).
24. F. G. Keyes, Thermal Conductivity of Gases, Trans. ASME 77, 1395-96 (1955).
25. F. G. Keyes and R. G. Vines, The Thermal Conductivity of Nitrogen and Argon, J. Heat Transfer 87, No. 2, 177-83 (1965).
26. A. Michels, J. V. Sengers, and L. J. M. Van de Klundert, The Thermal Conductivity of Argon at Elevated Densities, Physica 29, 149-60 (1963).
27. A. Uhlir, Thermal Conductivity of Fluid Argon and Nitrogen, J. Chem. Phys. 20, 463-72 (1952).
28. A. Eucken, Uber die Temperaturabhangigkeit der Warmeleitfahigkeit einiger Gase, (Concerning the Temperature Dependence of Thermal Conductivity of Several Gases), Physik Z. 12, 1101-17 (1911).
29. H. Ziebland, and J. T. A. Burton, The Thermal Conductivity of Nitrogen and Argon in the Liquid and Gaseous States, Brit. J. Appl. Phys. 9, 52-59 (1958).
30. L. D. Ikenberry and S. A. Rice, Kinetic Theory of Dense Fluids. XIV. Experimental and Theoretical Studies of Thermal Conductivity in Liquid Ar, Kr, Xe and CH₄, J. Chem. Phys. 39, 1561-71 (1963).
31. A. J. Rothman, Thermal Conductivity of Gases at High Temperatures, U. S. Atomic Energy Comm. UCRL 2339, 8-114 (1953).
32. W. G. Kannuliuk and E. H. Carman, The Thermal Conductivity of Rare Gases, Proc. Phys. Soc. (London) B65, 701-09 (1952).
33. F. Hutchinson, Self-Diffusion in Argon, J. Chem. Phys. 17, No. 11, 1081-86 (1949).
34. E. B. Winn, The Temperature Dependence of the Self-Diffusion Coefficients of Argon, Neon, Nitrogen, Oxygen, Carbon Dioxide and Methane, Phys. Rev. 80, No. 6, 1024- (1950).
35. I. Amdur and T. F. Schatzki, Diffusion Coefficients of the Systems Xe-Xe and A-Xe, J. Chem. Phys. 27, No. 5, 1049-54 (1957).
36. I. Amdur, Low Temperature Transport Properties of Gases. II. Neon, Argon, Krypton, and Xenon, J. Chem. Phys. 16, 190-94 (1948).
37. B. N. Srivastava and K. P. Srivastava, Mutual Diffusion of Pairs of Rare Gases at Different Temperatures, J. Chem. Phys. 30, No. 4, 984-90 (1959).
38. J. Kestin and W. Leidenfrost, An Absolute Determination of the Viscosity of Eleven Gases over a Range of Pressures, Physica 25, 1033-62 (1959).
39. J. Kestin and H. E. Wang, The Viscosity of Five Gases, A Re-evaluation, Trans. ASME 80, 11-17 (1958).
40. G. P. Flynn, R. V. Hanks, N. A. Lemaire, and J. Ross, The Viscosity of Nitrogen, Helium, Neon and Argon from -78.5 to 100°C below 200 Atmospheres, Brown University, Metclaf Chem. Lab., Providence, R. I., Tech. Rept. No. BRN-3-P (Nov 1962), NASA N63 11774.
41. M. Klein, The Determination of Intermolecular Potential Functions from Macroscopic Measurements, to be published.





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